

Formation of BH₃ Adducts with Pyridine-2-Methylaminophosphine ligands: An experimental and computational study

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Abstract. The reaction of pyridine-2-methylaminophosphine [C₅H₄N-CH₂NHPPH₂] (**1**) and pyridine-2-methylphosphinoselenoic amide [C₅H₄N-CH₂NHP(Se)Ph₂] (**2**) with BH₃·SMe₂ yields the corresponding adducts [C₅H₄N(BH₃)-CH₂NHP(BH₃)Ph₂] (**1a**), and [C₅H₄N(BH₃)-CH₂NHP(Se)Ph₂] (**2a**), respectively. The solid state structures of both the compounds were established by single crystal X-ray diffraction analysis. The phosphorus and the pyridine nitrogen atoms are coordinated to the boron atom in the case of **1a** whereas only pyridine nitrogen atom is attached to the BH₃ group in the case of **2a**. To understand the nature of P-B/ N-B bonds and to compare the basicities of pyridine nitrogen, amino nitrogen and phosphorus atoms, density functional theoretical (DFT) calculations were performed on the BH₃ adducts **1a** and **2a**. The results are consistent with the experimental results.

Keywords. Pyridine; Borane; DFT calculation; X-ray analysis; HOMO; LUMO.

1. Introduction

Various P-N ligands are often used in organometallic chemistry to achieve well-defined reaction centers.^{1,2} The basicity of the nitrogen atom adjacent to phosphorus in aminophosphines has been the subject of much discussion.³ Today it is well-accepted that in acyclic phosphinamines the tricoordinate nitrogen atom assumes a planar configuration with respect to its substituents and thus demonstrates diminished basicity due to enhanced N(pπ)-P(dπ) bonding.^{3,4} Due to lack of crystallinity, solid-state structural information on phosphinamine compounds in which phosphorus atom is trivalent is currently less compared to that for four-coordinate phosphorus compounds, which is inadequate for comparison. Previous studies have described that acyclic phosphinamines behave as "P-donors only" in their reactions with Lewis acid diborane.⁵ On the other hand, bis-(borane) adducts are well-known to form with phosphinamines of the type P(OCR₂CH₂)₂N (R = H, Me). Here, the nitrogen atom in the constrained bicyclic structure possesses a pyramidal geometry, which presumably weakens the pπ - dπ interaction and enables nitrogen to exhibit an enhanced basic character.⁶ Krannich *et al.*, reported by NMR studies, that even P-B bonded mono-adduct is thermodynamically preferable,

B-P-N-B bonded bis-adduct and N-B bonded adduct could also be realized by varying the mole ratio of boranes.⁷ However, structural evidences were not provided. Recently, we have undertaken a project to address the basicity of amido nitrogen attached to a phosphine group by both experimental work and theoretical calculations. In this regard we have reported various phosphinamines [Ph₂PNH(CHPh₂)], [Ph₂PNH(CPh₃)] and [Ph₂PNH(CMe₃)], [Ph₂PNH(2,6-Me₂C₆H₃)], and [Ph₂PNH(2,6-*i*-Pr₂C₆H₃)] which were introduced in alkali metal, alkaline earth metal and transition metal chemistry.^{8,9} We have demonstrated from the crystallographic analysis as well as DFT calculation that the P-B bonded mono-adducts are more stable than N-B bonded mono-adducts obtained from various phosphinamines ligands. In continuation with the previous study, we were interested to introduce a pyridyl functional group as a competing basic group with the aminophosphine ligand to compare the basicity of the pyridyl nitrogen, amino nitrogen and the phosphorus atom attached to the amino group towards BH₃.

Herein we report the synthetic and structural details of pyridine-2-methylamidophosphine borane bis-adduct [C₅H₄N(BH₃)-CH₂NHP(BH₃)Ph₂] (**1a**) and pyridine-2-methylphosphinoselenoic amide borane adduct [C₅H₄N(BH₃)-CH₂NHP(Se)Ph₂] (**2a**). We also describe the full accounts of DFT study for compounds **1a** and **2a** to compare the stability of the P-B bonded mono-adducts

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and N-B bonded mono-adducts or bis-adducts considering the energetic and electronic stability of the respective compounds.

2. Experimental

2.1 General Information

All manipulations of air-sensitive materials were performed with the rigorous exclusion of oxygen and moisture in flame-dried Schlenk-type glassware either on a dual manifold Schlenk line, interfaced to a high vacuum (10^{-4} Torr) line, or in an argon-filled M. Braun glove box. Hydrocarbon solvents (toluene and *n*-pentane) were distilled under nitrogen from LiAlH_4 and stored in the glove box. ^1H NMR (400 MHz), $^{13}\text{C}\{^1\text{H}\}$, $^{31}\text{P}\{^1\text{H}\}$ NMR (161.9 MHz) and $^{11}\text{B}\{^1\text{H}\}$ (128.4 MHz) spectra were recorded on a BRUKER AVANCE III-400 spectrometer. BRUKER ALPHA FT-IR spectrometer was used for FT-IR measurement. Elemental analyses were performed on a BRUKER EURO EA at the Indian Institute of Technology Hyderabad. Starting materials [$\text{C}_5\text{H}_4\text{N-CH}_2\text{NHPPH}_2$] (**1**), and [$\text{C}_5\text{H}_4\text{N-CH}_2\text{NHP(Se)Ph}_2$] (**2**) were prepared according to the literature procedures.⁸

2.2 Synthesis of [$\text{C}_5\text{H}_4\text{N}(\text{BH}_3)\text{CH}_2\text{NHP}(\text{BH}_3)\text{Ph}_2$](**1a**)

In a 25 mL flask, **1** (1.0 g, 3.41 mmol) was dissolved in toluene (10 mL), borane-dimethyl sulfide (0.64 mL, 6.82 mmol) was added drop wise onto this solution. The reaction mixture was then stirred for another 6 h. A white precipitate was formed and filtered off and dried under *vacuo*. Yield: 1.09 g. (99%). The title compound was re-crystallized from CH_2Cl_2 at -4°C . ^1H NMR (400 MHz, CDCl_3): δ 8.61 (d, $^3J_{\text{HH}} = 8.0$ Hz, 1H, ArH), 7.75-7.64 (m, 4H, ArH), 7.72 (d, 1H, $^3J_{\text{HH}} = 4.0$ Hz, Py-CH), 7.62 (d, 1H, $^3J_{\text{HH}} = 8.0$ Hz, Py-CH), 7.46-7.38 (m, 6H, ArH), 7.31-7.26 (m, 1H, Py-CH), 4.70-4.65 (dd, 2H, $J = 14.2$ Hz, 7.6 Hz, NH- CH_2 -Py), 4.25-4.21 (q, 1H, $J = 14.4$ Hz, 7.2 Hz, NH- CH_2 -Py), 2.60 (br, 3H, BH_3), 1.17 (br, 3H, BH_3) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, CDCl_3): δ 156.7 (PyC), 149.2 (PyC), 139.9 (PyC), 139.9 (P attached C), 134.3 (P attached *o*-ArC), 133.4 (P attached *p*-ArC), 128.5 (P attached *m*-ArC), 128.4 (P attached *m*-ArC), 127.0 (PyC), 123.84 (PyC), 44.5 (Py-C) ppm. $^{31}\text{P}\{^1\text{H}\}$ NMR (161.9 MHz, CDCl_3): δ 73.1 ppm. $^{11}\text{B}\{^1\text{H}\}$ NMR (128.4 MHz, CDCl_3): δ -14.3 ppm (N- BH_3), -38.4 ppm (P- BH_3). FT-IR (selected frequencies): $\nu = 3360$ (N-H), 2378 (B-H), 1432 (P-C), 998 (P-N), 824 (N-B), 601 (P-B) cm^{-1} .

2.3 Synthesis of [$\text{C}_5\text{H}_4\text{N}(\text{BH}_3)\text{CH}_2\text{NHP(Se)Ph}_2$](**2a**)

In a 25 mL flask, **2** (1.0 g, 2.68 mmol) was dissolved in toluene, borane-dimethyl sulfide (0.25 mL, 2.68 mmol) was added drop wise on to the solution. The reaction mixture was then stirred for another 6 h. A white precipitate was formed and was filtered through G4 frit and the white residue was dried under *vacuo*. Yield: 1.03 g (99%). The title compound was re-crystallized from CH_2Cl_2 at -4°C . ^1H NMR (400 MHz, CDCl_3): δ 8.66 (d, $^3J_{\text{HH}} = 8.0$ Hz, 1H, ArH), 7.84-7.74 (m, 4H, ArH), 7.73 (d, 1H, $^3J_{\text{HH}} = 4.0$ Hz, Py-CH), 7.62 (d, 1H, $^3J_{\text{HH}} = 8.0$ Hz, Py-CH), 7.46-7.38 (m, 6H, ArH), 7.31-7.26 (m, 1H, Py-CH), 4.75-4.70 (dd, 2H, $J = 14.2$ Hz, 7.6 Hz, NH- CH_2 -Py), 4.19-4.14 (q, 1H, $J = 14.4$ Hz, 7.2 Hz, NH- CH_2 -Py), 2.62 (br, 3H, BH_3) ppm. $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, CDCl_3): δ 156.7 (PyC), 149.2 (PyC), 139.9 (PyC), 139.9 (P attached C), 134.3 (P attached *o*-ArC), 133.4 (P attached *p*-ArC), 128.5 (P attached *m*-ArC), 128.4 (P attached *m*-ArC), 127.0 (PyC), 123.84 (PyC), 44.5 (Py-C) ppm. $^{31}\text{P}\{^1\text{H}\}$ NMR (161.9 MHz, CDCl_3): δ 57.6 ppm. $^{11}\text{B}\{^1\text{H}\}$ NMR (128.4 MHz, CDCl_3): δ -14.1 ppm (N- BH_3). FT-IR (selected frequencies): $\nu = 3270$ (N-H), 2379 (B-H), 1432 (P-C), 922 (P-N), 833 (N-B), 557 (P=Se) cm^{-1} .

2.4 Single-Crystal X-Ray Structure Determination and analysis

Single crystals of compounds **1** and **2** were obtained from the respective CH_2Cl_2 solution under inert atmosphere at a temperature of -4°C . In each case, a crystal of suitable dimensions was mounted on a CryoLoop (Hampton Research Corp.) with a layer of light mineral oil and placed in a nitrogen stream at 150(2) K. All measurements were made on an Agilent Supernova X-calibur Eos CCD detector with graphite-monochromatic Cu-K α (1.54184 Å) radiation. Crystal data and structure refinement parameters are summarized in table 1. The structures were solved by direct methods (SIR-2004)¹⁰ and refined on F^2 using the full-matrix least-squares method; using SHELXL-97.¹¹ Non-hydrogen atoms were anisotropically refined. H-atoms were included in the refinement on calculated positions riding on their carrier atoms. The function minimized was $[\sum w(F_o^2 - F_c^2)^2] (w = 1/\sigma^2(F_o^2) + (aP)^2 + bP)]$, where $P = (\text{Max}(F_o^2, 0) + 2F_c^2)/3$ with $\sigma^2(F_o^2)$ from counting statistics. The function R1 and wR2 were $(\sum ||F_o| - |F_c||) / \sum |F_o|$ and $[\sum w(F_o^2 - F_c^2)^2 / \sum (wF_o^4)]^{1/2}$, respectively. The ORTEP-3 program was used to draw the molecule. Crystallographic data (excluding structure factors) for the structures

Table 1. Crystallographic details of [C₅H₄N(BH₃)-CH₂NHP(BH₃)Ph₂] (**1a**) and [C₅H₄N(BH₃)-CH₂NHP(Se)Ph₂] (**2a**).

	1a	2a
CCDC No.	1418340	1418341
Formula	C ₁₈ H ₂₃ B ₂ N ₂ P	C ₁₈ H ₂₀ BN ₂ PSe
Formula weight	319.97	385.10
Crystal system	Monoclinic	Monoclinic
Space group	<i>P</i> 2 ₁ / <i>c</i>	<i>P</i> 2 ₁ / <i>c</i>
a, [Å]	13.6661(14)	13.9339(6)
b, [Å]	13.7280(17)	13.9605(5)
c, [Å]	9.9298(12)	9.6554(4)
α[°]	90	90
β, [°]	106.115(12)	106.272(4)
γ[°]	90	90
V, [Å ³]	1789.7(4)	1802.97(13)
Z	4	4
Density (mg/m ³)	1.188	1.419
T, [K]	150(2)	293(2)
Radiation	CuKα (λ = 1.54184 Å)	CuKα (λ = 1.54184 Å)
μ, mm ⁻¹	1.328	3.638
F(000)	680	784
Absorption correction	Multi-scan	Multi-scan
Reflections collected	7995	7879
Unique reflections	3374 [R _{int} = 0.0275]	3284 [R _{int} = 0.0468]
Completeness to Θ	97.8%	94.9 %
GOF	1.046	1.013
Refinement method	Full-matrix least-squares on F ²	Full-matrix least-squares on F ²
R ₁ ; wR ₂	0.0430; 0.1236	0.0470; 0.1180

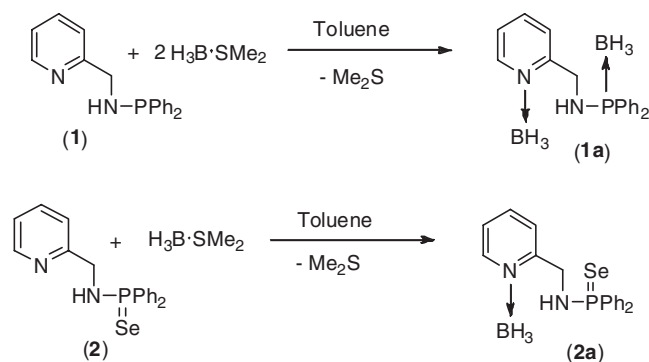
reported in this article have been deposited with the Cambridge Crystallographic Data Centre as supplementary publication nos. CCDC 1418340 (**1a**) and 1418341 (**2a**). Copies of the data can be obtained free of charge on application to CCDC, 12 Union Road, Cambridge CB21EZ, UK (fax: + (44)1223-336-033; email: deposit@ccdc.cam.ac.uk).

3. Results and Discussion

The bis-borane adduct of pyridine-2-methylaminophosphine [C₅H₄N(BH₃)-CH₂NHP(BH₃)Ph₂] (**1a**) was prepared in good yield by Lewis acid base reaction between the pyridine-2-methylaminophosphine (**1**) and BH₃·SMe₂ in 1:1 molar ratio in toluene at room temperature followed by recrystallization from dichloromethane solution (scheme 1). In an analogous reaction, the mono-borane adduct of pyridine-2-methylphosphinoselenoic amide [C₅H₄N(BH₃)-CH₂NHP(Se)Ph₂] (**2a**) was isolated by the reaction with pyridine-2-methylphosphinoselenoic amide (**2**) and H₃B·SMe₂ in 1:1 molar ratio (scheme 1). Attempts to isolate mono-borane adduct of **1** either [C₅H₄N(BH₃)-CH₂NHPPH₂] or [C₅H₄NCH₂NHP(BH₃)Ph₂] did not meet success as always compound **1a** was isolated as stable product. Both the

compounds **1a** and **2a** were characterized by spectroscopic/analytical techniques and the solid state structures of the compounds were established by single crystal X-ray diffraction analysis.

The formation of the bis-borane adduct **1a** from pyridine-2-methylaminophosphine and mono-borane adduct **2a** from pyridine-2-methylphosphinoselenoic amide can easily be followed by ¹H NMR spectroscopy, since additional resonances are observed for the borane (BH₃) groups attached to the pyridyl nitrogen and phosphorus atoms (for **1a**) and pyridyl nitrogen atom for

**Scheme 1.** Synthesis of borane-adducts **1a** and **2a** from compounds **1** and **2**, respectively.

2a as broad signals centered at δ 2.63 and 1.17 (**1a**), 2.62 (**2a**) ppm, respectively. The two methylene protons present in compounds **1** and **2**, resonate as doublet of a doublet centered at δ 4.67 (for **1a**) and 4.72 (for **2a**) due to coupling of two diastereotopic protons with adjacent N-H proton. As a result, two sets of coupling constants were observed for both the compounds (14.2 and 7.6 Hz). The pyridyl protons in compound **1a** display the resonances as two doublets centered at δ 7.73 and 7.62 ppm (δ 7.72 and 7.61 ppm for **2a**) and two multiplets at δ 7.46–7.38 and 7.31–7.26 ppm (δ 7.45–7.37 and 7.31–7.25 ppm for **2a**). In $^{13}\text{C}\{^1\text{H}\}$ NMR spectra of **1a** and **2a**, the resonances of pyridyl ring carbons were observed at higher chemical shift values (δ 156.7, 149.2 and 139.9 for **1a** and δ 156.5, 149.2 and 139.8 for **2a**) which are in agreement with the values reported in the literature.¹² The $^{31}\text{P}\{^1\text{H}\}$ of both the compounds show a sharp singlet at δ 73.1 and 57.6 ppm consistent with the presence of one phosphorus atom in each compound. However, the bis-borane compound is very sensitive towards oxygen and underwent oxidation to show an additional signal at δ 24.1 ppm which can be assigned to the P=O group (see Supporting Information). The $^{11}\text{B}\{^1\text{H}\}$ NMR spectra also confirmed the formation of bis-borane adduct for **1a** [δ -14.3 ppm, (PyN-BH₃), and -38.4 ppm, (P-BH₃)] and mono-borane adduct for **2a** (-14.1 ppm, PyN-BH₃). The $^{31}\text{P}\{^1\text{H}\}$ and $^{11}\text{B}\{^1\text{H}\}$ NMR data are very similar to our previously reported aminophosphine borane complexes of the type [Ph₂P(BH₃)NHR] (R = tBu, Ph₂CH, Ph₃C, 2,6-Me₂C₆H₃).⁸ In FT-IR spectra characteristic band for P-B bond stretching at 601 cm⁻¹ was observed for **1a** along with another absorption band at 2378 cm⁻¹ assigned for B-H stretching frequency. For compound **2a**, the absorption band at 557 cm⁻¹ can be assigned to P=Se bond along with a band at 2379 cm⁻¹ assigned to B-H stretching frequency. These values are in agreement with the values reported in the literature.¹³ To study in the solid state, the structures of the compounds **1a** and **2a** were established by single crystal X-ray diffraction analysis.

Both the bis-borane adduct **1a** and the mono-borane adduct **2a** crystallize in the monoclinic space group *P* 2₁/c having four independent molecules in the respective unit cell. The details of the structural parameters for compounds **1a** and **2a** are given in table 1. The molecular structures of **1a** and **2a** are shown in figures 1 and 2, respectively. The P–B bond distance of [1.909(2)] Å in **1a** is within the range of our previously reported values for amidophosphine borane compounds [Ph₂P(BH₃)NH(CR₃)] (R = Me, Ph, etc.) range.⁹ The B–N distances [1.597(2) Å for **1a** and 1.593(6) Å for **2a**] are also in the acceptable range to consider them

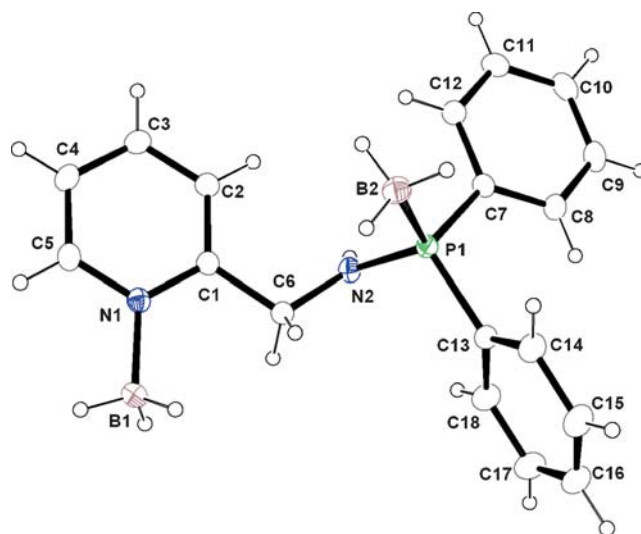


Figure 1. ORTEP diagram of **1a** with thermal displacement parameters drawn at the 30% probability level. Selected bond lengths [Å] and bond angles[°]: P1–N2 1.6669(13), P1–B2 1.9087(19), P1–C7 1.8099(16), P1–C13 1.8129(16), N1–C1 1.359(2), C1–C6 1.512(2), C6–N2 1.463(2), N1–B1 1.597(2), B2–P1–N2 112.08(8), B2–P1–C7 112.39(8), B2–P1–C13 112.54(9), P1–N2–C6 119.80(11), N1–C1–C6 117.10(14), B1–N1–C1 122.38(13), B1–N1–C5 118.37(14), N2–P1–C13 108.68(7), N2–P1–C7 103.64(7).

as coordinate bond. The P–N distances [1.6669(13) Å for **1a** and 1.657(3) Å for **2a**] are similar and indicate the presence of covalent bond between nitrogen and phosphorus atoms in both the compounds. Thus, in both the cases, the amino nitrogen is not basic enough to form an adduct with borane (BH₃) even if excess amount of borane is used. The solid state structures of all the aminophosphine-borane adducts are consistent with the dynamic nature of the molecules observed in multinuclear NMR spectra (*vide supra*).

Thus, we have observed that the basicities of pyridyl nitrogen and the phosphorus atoms in the aminophosphine moiety of compound **1** are almost similar and several attempts to isolate the mono-borane adduct of **1** either as [C₅H₄N(BH₃)-CH₂NHPPH₂] or [C₅H₄NCH₂NHP(BH₃)Ph₂] did not meet with success. In compound **2**, the phosphorus atom is blocked by a selenium functionality through the formation of P=Se bond and can not act as Lewis base as the P atom is now +5 oxidation state. Thus, the amino nitrogen (NH) and the pyridyl nitrogen atoms are now available for adduct formation with borane (BH₃). However, it is observed that only pyridyl nitrogen can form the adduct with BH₃ and the amino nitrogen (NH) remains inert even in the presence of excess amount of borane. Because of these observations, we were interested to calculate theoretically the limiting conditions for adduct formation

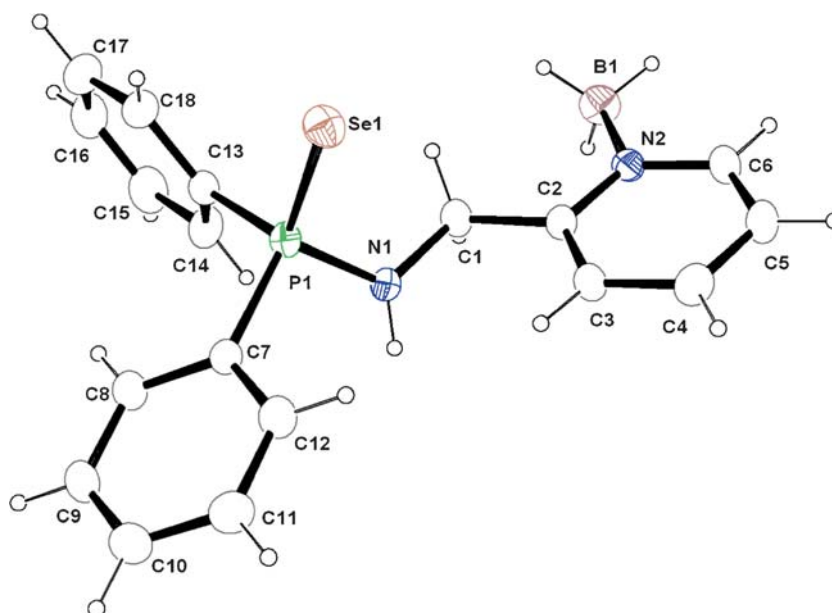


Figure 2. ORTEP diagram of **2a** with thermal displacement parameters drawn at the 30% probability level. Selected bond lengths [Å] and bond angles[°]: P1–N1 1.657(3), P1–Se1 2.1067(12), P1–C7 1.804(4), P1–C13 1.809(4), N1–C1 1.461(4), C1–C2 1.513(5), N2–B1 1.593(6), B1–N21–C2 123.2(4), B1–N2–C6 118.6(4), Se1–P1–N1 113.01(12), Se1–P1–C7 113.10(14), Se1–P1–C13 111.94(15), N1–P1–C7 102.49(18), N1–P1–C13 107.72(19), C7–P1–C13 107.99(19).

by pyridyl nitrogen, phosphorus atoms and the amino nitrogen using their frontier orbitals HOMO and LUMO.

3.1 DFT Calculations

3.1a General: Compound **1a** was obtained via the reaction of an equimolar ratio of the pyridine-2-methylaminophosphine ligand (**1**) and borane [BH₃·SMe₂]. The compound **1** contains three possible binding sites for the formation of borane adduct: (a) pyridyl nitrogen, (b) phosphorous, and (c) amino nitrogen. The first principle calculations were performed on the gas phase addition of borane (BH₃) to various pyridine-2-methylaminophosphine ligands. The main focus of our study was to explore the preferential Lewis basicity of pyridyl nitrogen, amino nitrogen or phosphorous atoms (depending on the ligand) upon reaction with borane

(BH₃). We have chosen several model ligand systems (figure 3 **a-e**) for this purpose along with experimentally synthesized compounds (**1a** and **2a**). All the model structures were drawn in Gauss View¹⁴ and optimized at HF/3-21G(d) level. Higher level single point energy calculations were done at B3LYP/6-311+G(2d,p)¹⁵ level of electronic structure calculation using Gaussian 09(Rev B.01) software.¹⁶ We have computed the structure and energetics of each of the free ligands listed in figure 3, the free BH₃ molecule, and each pyridine-2-methylaminophosphine-borane compounds by following the geometry optimization procedure. The optimized structures were confirmed to be (local) minima by the absence of imaginary vibrational frequencies. Theoretical analysis of all the systems for the adduct formation with borane (BH₃) was presented considering the stabilization energy of each adduct and subsequently, further investigations have been carried out in

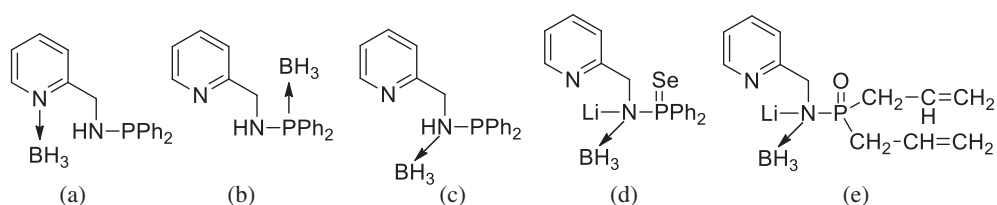


Figure 3. Selected pyridine-2-methylaminophosphine ligands and its chalcogen derivatives used for DFT calculations.

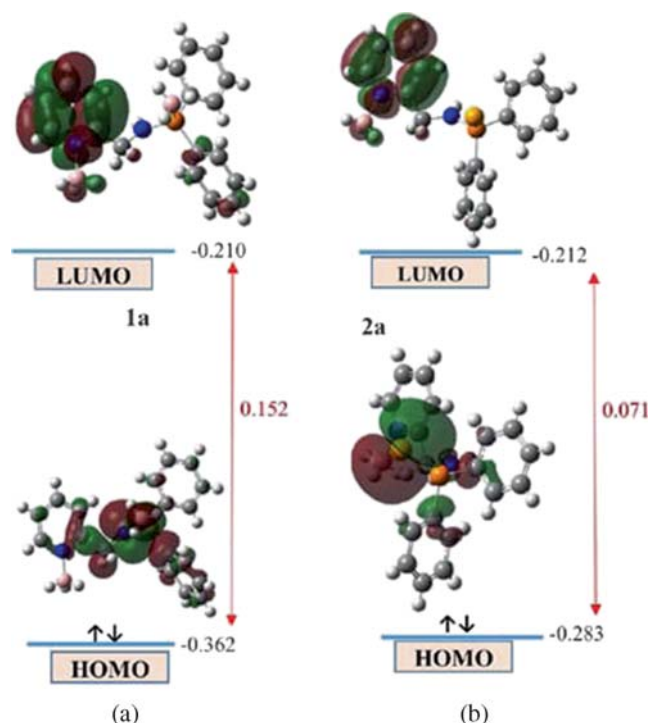


Figure 4. Frontier KS-molecular orbitals for bis-adduct of **1a** and mono-adduct of **2a**.

the quest for the reaction center of BH_3 at the less basic amino nitrogen site. After a random search considering various electronic and steric factors through different functional groups (results are not shown), the ligands **3d** and **3e** were found to be appropriate for formation of mono-adducts involving amino nitrogen. We have also performed frontier KS-molecular orbital calculations to investigate the electronic stability of adducts.

3.1b DFT analysis: The reaction of BH_3 with ligand (**1**) for the formation of bis-adduct (**1a**) is exothermic with stabilization energy of -64.7 kJ/mol. The phosphorus, pyridyl and amine nitrogen atoms bonded mono-adducts are stabilized by -33.2 , -30.6 and -22.5 kJ/mol, respectively. The values indicate the least basicity for the amine nitrogen bonded adduct; other two reaction sites have significant values of binding energies for mono-adducts. Our predicted stability for bis-adduct complements the experimental results regarding the possible formation of **1a** over the other theoretically obtained mono-adducts. The mono-borane adduct (**2a**) having blocked at P-site, has two possible sites for reaction with BH_3 ; pyridyl and amine nitrogen atoms. However, the adduct formed through former type nitrogen is stable by -28.5 kJ/mol as compared to the latter. After studying many functional groups in the quest of amine nitrogen bonded adducts, we have found that mono lithiated compounds (P blocked) having phenyl and vinyl functional groups attached to the phosphorus atom can serve the purpose of stabilizing amino nitrogen bonded mono borane-adducts. The suitable choices of p-substituted functional groups helped us to tune the basicity of two nitrogen sites. The complexes **3d** and **3e** have stabilization energies of values -30.7 and -43.5 kJ/mol, respectively. The adducts formed by the reactions of BH_3 to the pyridyl nitrogen centers of ligands, **3d** and **3e**, have stabilization energies of values -27.9 and -39.0 kJ/mol, respectively.

We also calculated the HOMO–LUMO energy gap for adducts, and the orbital pictures are shown in figures 4 and 5. The molecular orbital analysis is also

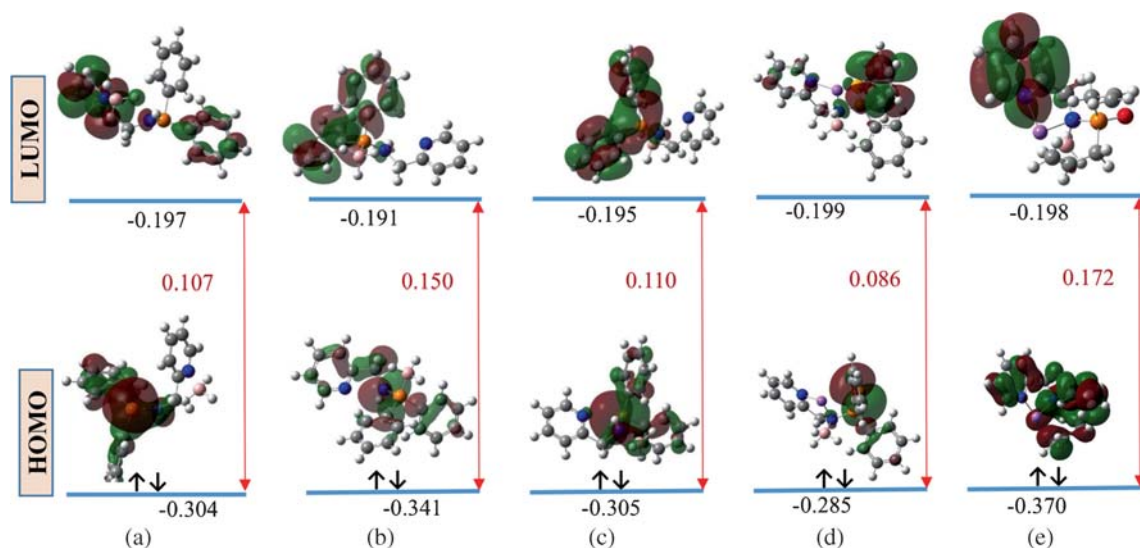


Figure 5. Frontier KS-molecular orbitals for mono-borane adduct of **1** (**a**, **b** and **c**), lithium (divinylphosphinoxide-2-pyridinylmethyl) ligand (**d**) and lithium (diphenylphosphinoselenoyl-2-pyridinylmethyl) (**e**).

well-consistent with the energetic investigations. From the HOMO–LUMO calculations, it is evident that the bis-adduct is more favorable compared to other mono-borane adducts; the energy gap (0.152 eV) of bis-adduct **1a** is higher than other mono borane adducts (**3a**, **3b** and **3c**). As a consequence, we could isolate only bis-adduct of pyridine-2-methylaminophosphine experimentally.

When we used penta-valent phosphorus atom by blocking with selenium atom in compound **2** (scheme 1), we observed that only pyridyl nitrogen showed tendency for adduct formation, and amino nitrogen (-NH) remained unactivated even in the presence of excess amount of borane. A further interesting observation is that blocking at phosphorus atom with Se atom results in the destabilization of the product **2a** (0.071 eV) with respect to the bis-borane adduct, **1a** (0.152 eV). In the phosphorus(V) atom, due to N(p π) \rightarrow P(d π) back bonding, the electron density of the amino nitrogen is reduced to some extent in comparison with pyridyl nitrogen which remains free. The lone pair of electrons resides outside of the plane of the ring and does not participate in aromatic delocalization. So in case of compound **2a**, favorable adduct formation is at the pyridyl N site. One of our goals was to activate the less basic amino nitrogen site. From our computational study, we observed that the accumulation of free negative charge at amino nitrogen center via the formation of lithiated diphenylphosphinoselenoyl-2-pyridinylmethyl (**3d**) and lithiated divinylphosphinooxide-2-pyridinylmethyl (**3e**) compounds are able to make σ -donation to the electron deficient borane. The electron cloud is more localized at amino nitrogen to behave as Lewis base. The latter ligand, having HOMO-LUMO gap of 0.172 eV has higher potential to bind with BH₃ through amino nitrogen atom. Computationally, we also have found out a way to activate the less basic amino nitrogen site and as it can be seen that this mono-borane adduct is more stabilized by nearly -0.022 eV than the corresponding bis-adduct **1a**. This mono-borane amine adduct can be useful as a polymer stabilizer if it is experimentally realized.

4. Conclusions

In summary, we have presented the synthesis and structural characterization of pyridine-2-methylaminophosphine-borane adducts, **1a** and **2a**; formation of the bis-borane adduct (**1a**) indicates similar basicity of the pyridyl nitrogen and the phosphorus atoms in pyridine-2-methylaminophosphine. However, the mono-borane adduct **2a** is formed through coordination of pyridyl

nitrogen when phosphorus atom is blocked by selenium atom via the formation of P=Se. In addition, we have also demonstrated that the P-coordinated borane is more stable to form N-coordinated borane by DFT calculations. However, the N-coordinated borane adduct resulting from the pyridine nitrogen can be stabilized if the phosphorus atom is blocked with chalcogenides. We have also theoretically observed that less basic amino nitrogen site can also have enhanced basicity through the formation of lithium complex of the corresponding ligands.

Supplementary Information

The NMR spectra of compounds **1a** and **2a** and all the theoretical calculation details are given in supplementary information which is available at www.ias.ac.in/chemsci.

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