

Synthesis and characterization of submicron-sized mesoporous aluminosilicate spheres

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MS received 27 March 2001

Abstract. Mesoporous aluminosilicate spheres of 0.3–0.4 μm diameter, with different Si/Al ratios, have been prepared by surfactant templating. Surface area of these materials is in the 510–970 $\text{m}^2 \text{g}^{-1}$ range and pore diameter in the 15–20 \AA range.

Keywords. Mesoporous solids; mesoporous aluminosilicate spheres; surfactant templating.

1. Introduction

The discovery of mesoporous silica marks a major event in areas related to catalysis, sorption and separation processes¹. Amongst the criteria for the use of such mesoporous materials for various purposes, morphology and particle size are important ones. In this context, several workers have attempted to prepare submicron sized mesoporous silica spheres^{2,3}. The synthetic procedure generally makes use of Stöber's reaction involving base hydrolysis and subsequent condensation of the silicon ester in an appropriate solvent⁴. One of the first efforts to make spherical silica is due to Unger *et al*⁵, who obtained non-porous silica beads with a particle size of $\sim 3 \mu\text{m}$. Kaiser and Unger⁶, however, obtained completely porous silica spheres by a similar procedure. Büchel *et al*⁷ have described a novel pathway for obtaining monodisperse silica spheres composed of a non-porous solid silica core and a thin mesoporous silica shell by taking advantage of the methods described by Unger *et al*⁵ and Kaiser and Unger⁶. By employing surfactant-stabilized emulsions as in the synthesis of MCM-41, Huo *et al*³ obtained large, monodisperse mesoporous silica spheres with high surface areas. The spheres had diameters between 0.1 to 2 μm and pore diameters of 1 to 5 nm. Mesoporous silica spheres ranging from 2 to 6 μm (pore diameter ~ 2.5 nm) have been synthesized under static acidic conditions by mixed cationic-nonionic surfactant templating as well⁸. Spheres of ordered MCM-41 and MCM-48 have also been prepared by employing appropriate surfactants^{9,10}. It was our interest to

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prepare monodisperse aluminosilicate spheres with mesoporous properties. We have been able to prepare such spherical aluminosilicates with nominal Si/Al ratios between 14 and 40 by employing a cationic surfactant. For the purpose of comparison, we also prepared mesoporous silica spheres by the same procedure.

2. Experimental

2.1 Synthesis

We have prepared mesoporous aluminosilicate spheres by the following procedure. In a typical synthesis, 0.667 g N-cetyl-N,N,N-trimethylammonium bromide (CTAB) was dissolved in a solution containing 14 mL deionised water (H_2O) and 14 mL ethanol ($\text{C}_2\text{H}_5\text{OH}$). To the above solution, 3.5 mL ammonia (NH_3) solution (25%) was added under stirring, followed by the addition of a solution of 15 mL of an aqueous solution containing 0.1 g $\text{Al}_2(\text{SO}_4)_3$. The final composition of the mixture was 0.4 M CTAB/360 M H_2O /53 M $\text{C}_2\text{H}_5\text{OH}$ /41.8 M NH_3 /1 M TEOS/0.04M $\text{Al}_2(\text{SO}_4)_3$. After 10 min, 1 mL of tetraethylorthosilicate (TEOS) was added dropwise over a period of 2 h under stirring to obtain a gel. The gel was aged at room temperature (298 K) for 16 h, filtered, washed with deionised water and dried at ambient temperature. The dried product was calcined at 200°C for 4 h, which resulted in sample C. By a similar procedure, we prepared two other aluminosilicate samples (A and B) with different Si/Al ratios. Si/Al ratios in the starting compositions were 14, 24 and 40 respectively in samples A, B and C. Sample D corresponds to pure silica.

2.2 Characterization

Powder X-ray diffraction (XRD) patterns of the samples were obtained using $\text{CuK}\alpha$ radiation (Rich-Siefert, XRD-3000-TT). Thermogravimetric analysis (TGA) was carried out (Mettler Toledo, TG-850) in flowing oxygen atmosphere (50 mL min^{-1}) at a heating rate of 10 K min^{-1} . Scanning electron microscopy (SEM) and energy-dispersive X-ray (EDX) analysis were carried out using a Leica scanning electron microscope fitted with EDX spectrometer and a Link ISIS detector. Surface areas were determined by the Brunauer-Emmet-Teller (BET) method on a Belsorp 28 instrument. Pore-size distribution was calculated by the Barrett-Joyner-Halenda (BJH) method¹¹. Solid state MAS NMR measurements were performed on a Bruker DSX-300 MHz Spectrometer. Tetramethylsilane (TMS) and $\text{Al}(\text{H}_2\text{O})_6^{3+}$ were used as external standards.

3. Results and discussion

Table 1 lists the compositions of the various samples prepared by us. Calcination of the spherical aluminosilicates (A–C) and silica (D) at 200°C for 4 h removed most of the surfactant template as will be demonstrated later. In figures 1 and 2, we show the SEM images of the two aluminosilicate samples (as-synthesized and calcined) with Si/Al ratios of 14 and 40 (A and C respectively). The as-synthesized sample of A with a Si/Al ratio of 14 shows spherical morphology as revealed by the SEM images. There is little agglomeration of particles. The spherical

Table 1. Properties of mesoporous aluminosilicate spheres.

Sample	Si/Al ^a	d_{100} ^b (Å)	Particle diameter (µm)	Surface area (m ² g ⁻¹)	Pore diameter (Å)
A	14 (49)	35.8 (36.2)	0.35	970	18
B	24 (82)	36.8 (35.3)	0.40	510	15
C	40 (132)	34.8 (30.6)	0.30	700	19
D	Infinity	31.3 (29.6)	0.43	1100	20

^aValues in parentheses are obtained after taking into account Al in the framework;
^b d -value of the low angle reflection in the XRD pattern. Values in parentheses were obtained after calcination at 200°C for 4 h. After calcination at 473 K for 4 h, the percentage of template removed in the samples is 87–95%

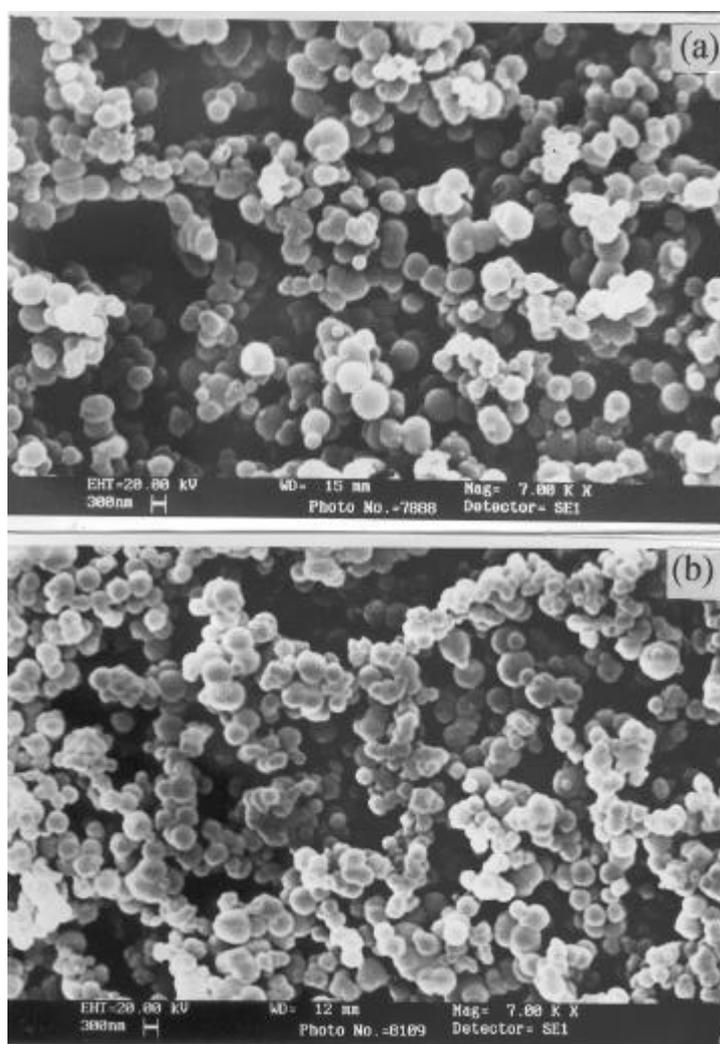


Figure 1. SEM micrographs of aluminosilicate spheres (A): (a) as-synthesized and (b) after calcination at 200°C for 4 h.

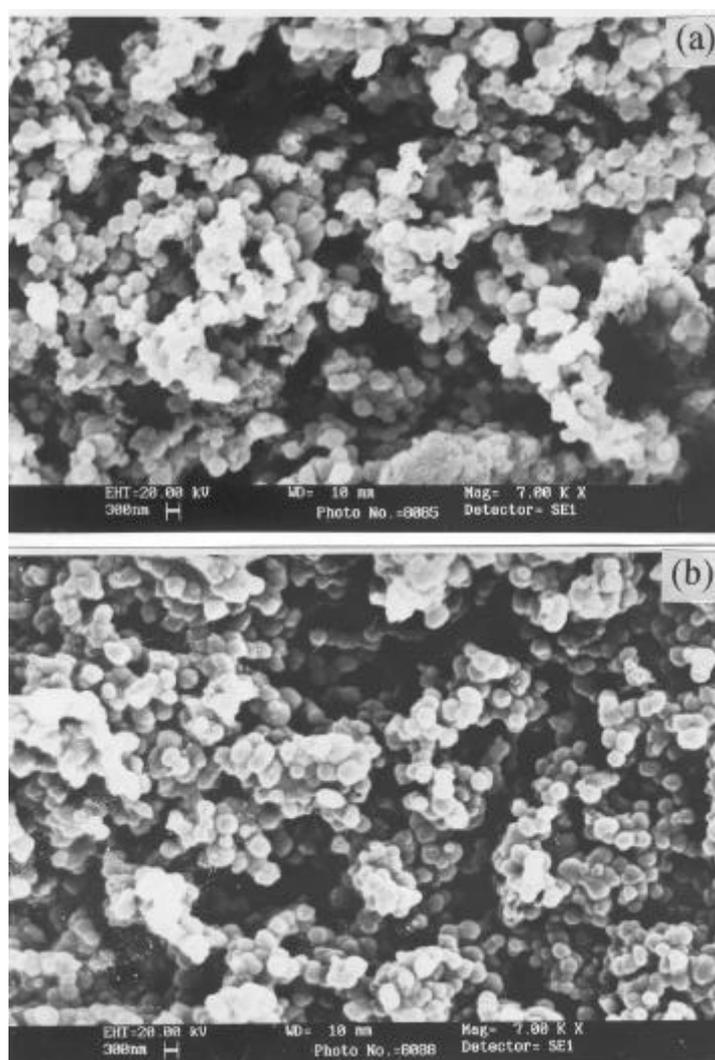


Figure 2. SEM micrographs of aluminosilicate spheres (C): (a) as-synthesized and (b) after calcination at 200°C for 4 h.

morphology is retained even after calcination at 200°C, as can be seen from figure 1b. This was true of the other two aluminosilicate samples (B and C) as well, as demonstrated in figure 2b in the case of sample C. The particles show a narrow size distribution with average diameters of 0.35 and 0.30 μm in the case of A and C respectively. The article sizes of all the samples are listed in table 1.

Powder XRD patterns of the as-synthesized aluminosilicate and silica spheres are presented in figure 3a. The XRD patterns show a single peak at 2θ values of 2.46°, 2.38°, 2.54° and 2.81° respectively for A, B, C and D. The low-angle peaks are similar to those of the hexagonal mesoporous solids. The powder XRD patterns of the samples after calcination at 200°C for 4 h exhibit single broad peaks at 2θ

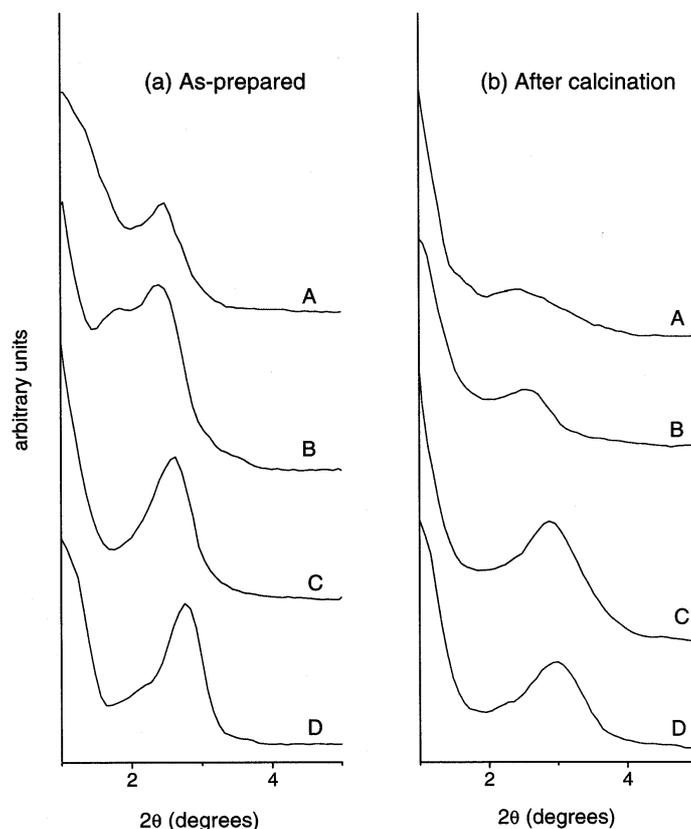


Figure 3. XRD patterns of the aluminosilicate (A, B and C) and pure silica (D) spheres: (a) as-synthesized and (b) after calcination.

values of 2.44° , 2.5° , 2.88° and 2.98° respectively for A, B, C and D as shown in Figure 3b. The shift in the peak positions after calcination is due to the removal of the template. The presence of the low-angle reflection after calcination indicates that the mesophase is reasonably intact even after the removal of the template, although the intensity is considerably diminished in the case of the samples A and B. The spherical morphology is, however, retained as mentioned earlier. It appears that Al-rich samples are more disordered.

TGA curves of the aluminosilicates A and C are shown in figure 4. The as-synthesized samples show one major mass loss in the temperature range $200^\circ\text{--}300^\circ\text{C}$. The mass loss is around 40%, where most of the template is removed. Thus, the sample calcined at 200°C for 4 h, only shows a mass loss ($< 20\%$) due to adsorbed water around 100°C . The TGA curves confirm that most of the template is removed on calcination at 200°C for 4 h. These observations are consistent with those of Tanev *et al*¹², who have reported similar TGA curves for mesoporous silica before and after calcination. We did not calcine the samples at temperatures higher than 200°C in order to ensure that the morphology and surface area of the

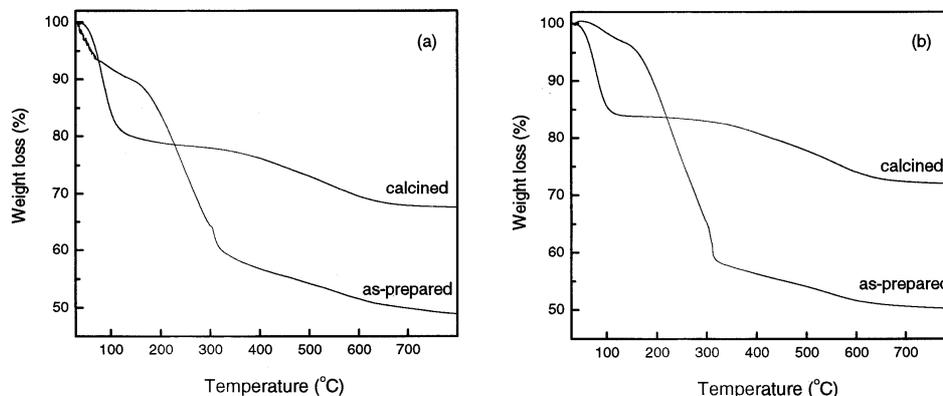


Figure 4. TGA curves of the aluminosilicate samples A (a) and C (b).

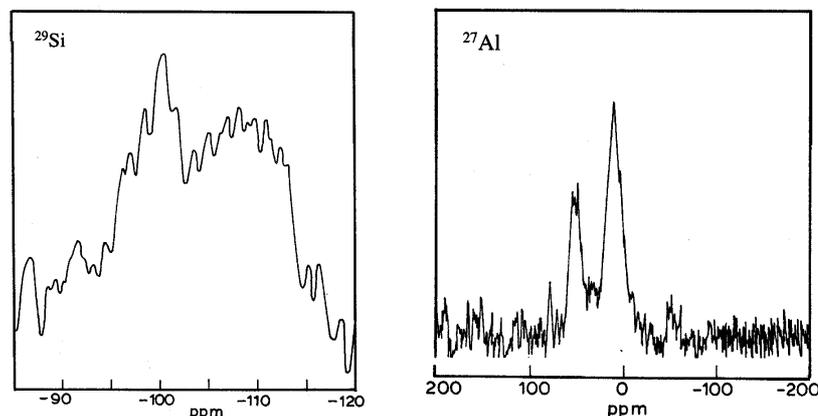


Figure 5. ^{29}Si and ^{27}Al MAS-NMR spectra of as-synthesized sample C.

samples are not badly affected. Percentage values of the template removed on calcination at 200°C are in the range 87–95%.

^{29}Si and ^{27}Al MAS NMR measurements were carried out on the as-synthesized and calcined sample C. In figure 5 we show the ^{29}Si and ^{27}Al MAS NMR spectra of the as-synthesized sample C. The ^{29}Si NMR spectrum of the as-synthesized sample shows two peaks at -100 and -108 ppm due to the Q^3 and Q^4 silicate species respectively¹³. The intensity of the Q^3 species is greater than that of the Q^4 species indicating that the silanol groups have not condensed fully. On calcination, however, the relative intensity of the Q^4 species increases, showing that the silanol groups condense to form the Si-O-Si species. The ^{27}Al NMR spectrum of the as-synthesized samples show two peaks at 47 and 9.5 ppm. These are due to tetrahedrally and octahedrally co-ordinated species respectively¹⁴. After calcination, the relative proportion of the peaks corresponding to the tetrahedral and the octahedral species remains the same. Taking account of only the tetrahedral

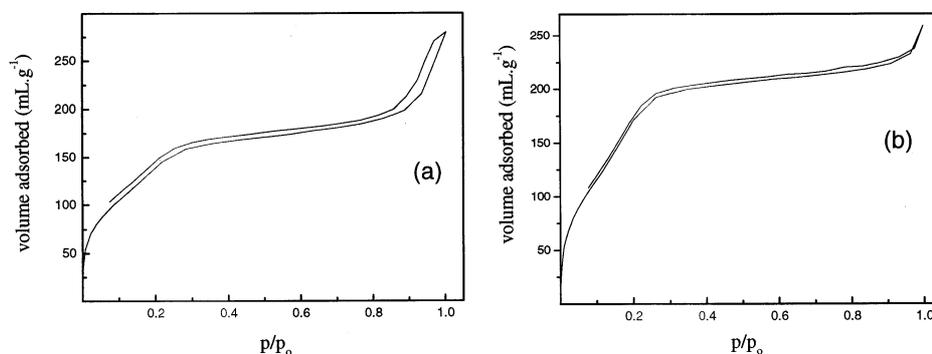


Figure 6. N_2 adsorption isotherms of (a) sample A calcined at 200°C for 4 h, and (b) sample C calcined at 200°C for 4 h.

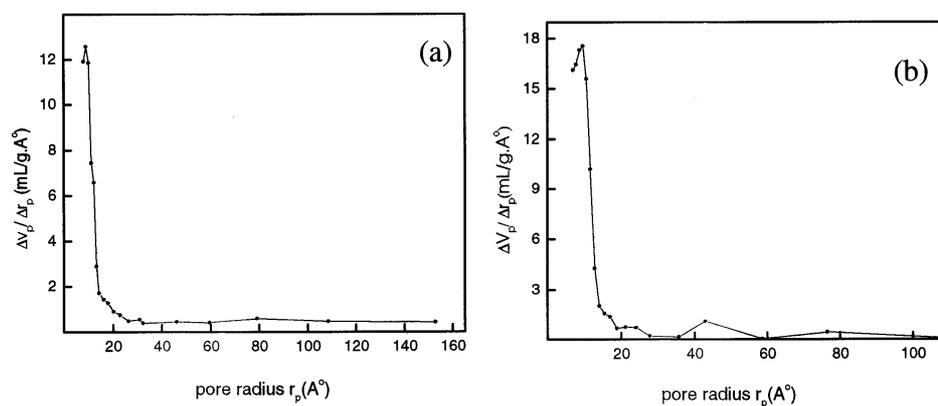


Figure 7. Pore size distribution curves of the calcined samples A (a) and C (b) obtained by using the BJH method.

species as a part of the framework, we estimate the percentage of aluminium in the framework to be around 30%. The Si/Al ratios calculated on this basis are listed in table 1.

Nitrogen adsorption isotherms of samples A and C are presented in figure 6. These are typical type IV adsorption isotherms. The BET surface areas of the samples calculated from N_2 adsorption are in the range $500\text{--}1000\text{ m}^2\text{ g}^{-1}$ (table 1). In general, the surface area of the aluminosilicate spheres is smaller than that of pure silica spheres. Surface area of the pure silica spheres as determined by us is $1100\text{ m}^2\text{ g}^{-1}$, the literature⁹ value being $1200\text{ m}^2\text{ g}^{-1}$. Pore size distribution was calculated using the BJH method. In figure 7, we show the plots of the pore size in case of samples A and C. The particles have a narrow pore size distribution, the values of the pore diameters being in the range $15\text{--}20\text{ \AA}$ (table 1). The pore size of the aluminosilicate spheres is smaller than that of the pure silica spheres. We could estimate the diameter of the pores from the low-angle XRD peaks by making use of the relation $a_0 = (2d_{100})/(3)^{1/2}$. The values of the diameter obtained are considerably

larger than the pore diameter from adsorption measurements, since the former includes the wall thickness. We have estimated the wall thickness from the difference in the XRD (a_0) and adsorption pore diameters. These values are in the 7–13 Å range, the thickness increasing with Al content.

4. Conclusions

In summary, the present work demonstrates that it is possible to prepare submicron-sized mesoporous spheres of aluminosilicates by using surfactant templating. Although the sample crystallinity is not as high as in the case of pure silica spheres prepared by a similar procedure, the high surface areas and desirable pore diameters of the aluminosilicate spheres make them good candidates for use in catalysis, sorption and related applications.

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