

Charge-transfer spectra of ferrocene in halocarbon solvents under photoexcitation

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Abstract. The changes in the electronic absorption spectra of ferrocene in the halocarbon solvents chloroform and carbontetrachloride have been investigated under photoexcitation in nitrogen atmosphere. Photoexcitations have been made with monochromatic light (using an Xe-source and a monochromator), at intervals of a few nanometers in the spectral range 220–750 nm. Analysing the spectra by a modified method the position of the charge-transfer-to-solvent (CTTS) band has been located for both the solvents. The position of the CTTS band in the case of carbontetrachloride solution located (320 nm) by the present study is different from the previously reported value (307 nm), while from the previous studies the position of the CTTS band in the case of the spectra of ferrocene in chloroform was not clear. From the present investigation, the changes in spectra after photoexcitation studied as a function, the concentration of ferrocene in the solution and the time (duration) of photoexcitations, have been observed to be systematic. Using the position of the new band (320 nm) for the CTTS transition in the case of carbontetrachloride, ionization potential of ferrocene has been estimated and the estimated value has shown excellent agreement with the experimental value indicating the exactness of the newly located CTTS band position.

Keywords. Ferrocene; charge-transfer; halocarbon solvents; photoexcitation.

1. Introduction

Ferrocene (*bis*(cyclopentadienyl) iron: $(C_5H_5)_2Fe$; abbreviated as FcH, the first synthesized metallocene) has wide application in various technological fields¹. The important electrochemical, photochemical and photophysical properties of the material have been recognized^{2–5}. This material has been used (i) efficiently as mediators in various electron transfer processes², (ii) in the development of biosensors² and (iii) for the synthesis⁶ of new materials of higher electrical conductivity from poly (vinylidene chloride). Recent studies on the semi- and photoconductivity of this material under various experimental conditions have shown very interesting results^{7–9}. Literature survey has shown that the studies on the change in electronic absorption spectra of ferrocene after photoexcitation (at several wavelengths in the UV region) in halocarbon solvents have received much attention in the past^{10–13}. Formation of charge-transfer complexes of ferrocene with halocarbon solvents after photoexcitation (ferrocene- to solvent charge-

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transfer (CTTS)) has been predicted by several research workers¹⁰⁻¹³. From the previous study¹² the position of the CTTS band in the case of CCl₄ solution was reported at 307 nm but the position of the CTTS band in the case of CHCl₃ was not clear. The dissociation of the CTTS excited state to ferricenium cation was evidenced^{11,12} by the characteristic absorption band at \approx 617 nm. However, the earlier studies could not properly locate the position of the absorption band arising from the CTTS transition in the case of solutions of ferrocene in both CCl₄ and CHCl₃. This requires further studies on the topic. Study on the charge-transfer-to-solvent photochemistry are also a topic of current research as evidenced by the work of Tatistcheff *et al*¹⁴.

Previously, Brand and Snedden¹² using the wavelength of 365 nm for excitation studied the effect of photoexcitation on the spectra of ferrocene in CCl₄. Traverso and Scandola¹¹ used several wavelengths (313, 334, 365, 404 and 472 nm) for photoexcitation to study the effect of photoexcitation on the spectra of ferrocene in CCl₄ and CHCl₃ and they concluded that the irradiation in the spectral region where the ferrocene-to-halocarbon solvent charge transfer bands overlap the intramolecular bands caused the oxidation of ferrocene to ferricenium cation. From the previous limited experiment on photoexcitation the exact profile of the action spectrum for the photoinduced changes could not be obtained. Therefore, attempts were made to find out the exact profile of the action spectrum for the photoinduced changes using excitation wavelength at intervals of a few nanometers in the spectral range 220–750 nm.

In the present study, the effects of photoexcitations (using monochromatic light) on the spectral behaviour have been studied in the nitrogen environment as a function of different parameters like exciting wavelength (at intervals of a few nanometers in the spectral range 220–750 nm), the concentration of ferrocene in the solution and the time (duration) of light exposure. In the previous studies the spectra obtained after photoexcitations were analysed directly but in the present study instead of analysing the spectra obtained after photoexcitations, the difference in spectra expressed by the change in absorbance (i.e. absorbance after photoexcitation minus absorbance before photoexcitation) have been analysed to locate properly the positions of the new absorption bands. Analysing the spectra by the modified method as mentioned above, exact position of the charge-transfer-to-solvent (CTTS) band has been located for both the solvents. Using the new band position for CTTS transition in the case of carbontetrachloride, ionization potential of ferrocene has been estimated and the estimated value has shown excellent agreement with the experimental value indicating the exactness of the newly located CTTS band position. The results are presented in this article.

2. Experimental

High purity microcrystalline ferrocene in powder form was obtained from Strem Chemicals (Newburyport, MA, USA) and was used after further purification by repeated crystallization. Solvents used e.g., carbontetrachloride and chloroform were of spectrograde quality. The purity of the material was checked by usual spectral investigation.

Solution of ferrocene was made in the above-mentioned solvents. The irradiations were performed using 1 cm spectrophotometer cells filled with \approx 3 ml of solution at \approx 300 K and using monochromatic light of various wavelengths in the range 220–750 nm. Fresh solution was used for each experiment. For the studies on

photoexcitations, the solutions were deoxygenated by saturation with pure and dry nitrogen gas. For removing the dissolved oxygen from the solution, dry nitrogen gas was allowed to bubble through (for about 15 min; using glass capillary) the sample solution kept in the quartz cell before photoexcitation and then photoexcitation was carried out under the presence of nitrogen gas flow also. In our experimental conditions in pure solvents no clouding of solution was observed for at least initial one hour after photoexcitation. However, if the solution was kept for a long time (several hours) a little precipitation was noticed at the bottom of the sample cell. The electronic absorption spectra were recorded before and after the photoexcitation by using a UV-Vis Scanning Spectrophotometer, Model UV-2101 PC (Shimadzu, Japan). The resolution of the monochromator of this spectrophotometer is 0.1 nm. A Xe-source operated at 60 W (Spectral Energy, USA) and a monochromator (type-H-20 UV, Ins. SA, division: Jobin Yvon, France) were used for obtaining radiation of different monochromatic wavelengths. The error limit involved in the excitation wavelength is ~ 0.2 nm and the slit width used during photoexcitation was 2 mm for all the cases. For the excitation with visible light, a glass filter was used in the path of the excitation beam before the monochromator to eliminate the second order diffraction light (UV) from the monochromatic light in the visible range. A Luxmeter (Model: 5011, Aplab, India) was used to estimate the photon flux of the excitation beam reaching the sample solution. Maximum photon flux (corresponding to $\lambda = 290$ nm) reaching the sample cell was $\sim 8.6 \times 10^{16}$ quanta/m²/s.

3. Results and discussion

Solutions of ferrocene in CHCl₃ or CCl₄ in the dark are stable for long periods and this phenomenon has been identified by spectral analysis. It has been reported^{3,11} that if solutions of ferrocene in CHCl₃ or CCl₄ are irradiated with UV radiation of proper wavelength, photooxidation of ferrocene occurs. We have studied systematically the effects of irradiation of both ultraviolet and visible light. The electronic absorption spectra of ferrocene in CHCl₃ and CCl₄ solvents before as well as after photoexcitation with UV light are shown in figure 1. For the solution of ferrocene in CHCl₃, the electronic absorption spectrum before photoexcitation (curve 1, figure 1) shows a broad band at ≈ 440 nm and in this case the appearance of the band in the UV region is not clear but there is an indication of a very weak shoulder at ≈ 325 nm. The effects of the irradiation (5 min) of UV light, at 285 nm, on the solution of ferrocene in CHCl₃ are shown in curve 2, figure 1. Sample solution in CCl₄ before photoexcitation shows (curve 3, figure 1) one broad band at ≈ 440 nm and another band at ≈ 307 nm. The effects of the irradiation of UV light (10 min), at 320 nm, on the solution of ferrocene in CCl₄ are shown in curve 4, figure 1. From figure 1 it is seen that intensity of the spectra after photoexcitation increases significantly in the UV region, compared to the visible region. In addition, appearance of a new band at longer wavelength side (at ≈ 620 nm) is clear from figure 1. In figure 2 the difference in spectra, expressed by the change in absorbance (i.e. absorbance after photoexcitation minus absorbance before photoexcitation) as a function of wavelength is shown for both the solvents. From figure 2, appearance of new peaks and shoulders in the spectrum for both the solvents are clear. Thus, figure 2 demonstrates clearly the effects of irradiation of UV light upon the solution of CHCl₃ and CCl₄ on the basis of spectral changes. From curve 1 in figure 2 it is observed that the spectrum for CHCl₃ shows three new sharp absorption bands at ≈ 290 , 362 and 620 nm

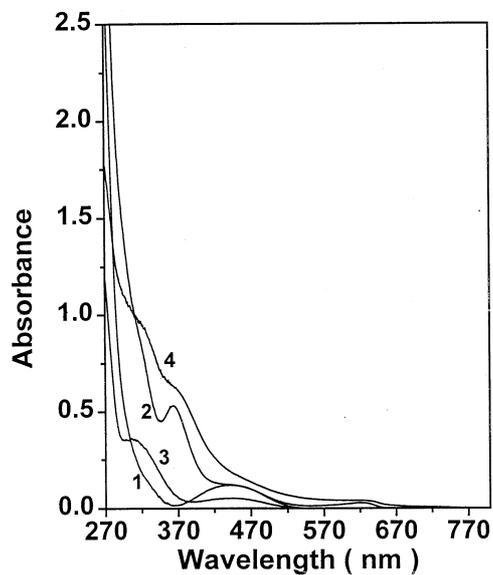


Figure 1. Electronic absorption spectrum of ferrocene in CHCl_3 and CCl_4 solutions before and after photoexcitation: 1, in CHCl_3 (concentration, 6.8×10^{-4} M) solution before photoexcitation; 2, in CHCl_3 solution after photoexcitation by 285 nm for 5 min; 3, in CCl_4 (concentration 2.04×10^{-4} M) solution before photoexcitation and 4, in CCl_4 solution after photoexcitation by 320 nm for 10 min.

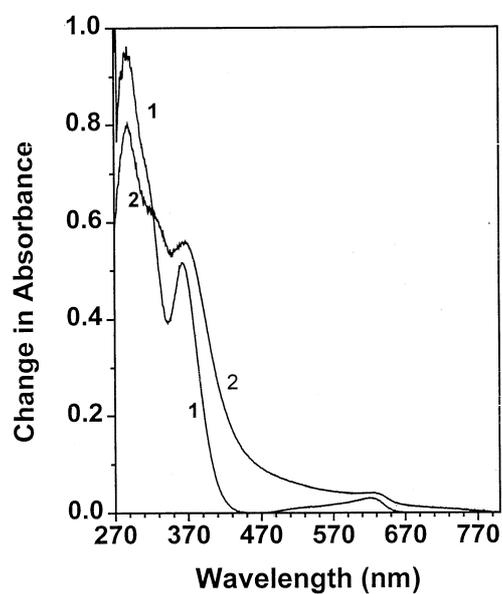


Figure 2. Change in absorbance versus wavelength for solutions of ferrocene in CHCl_3 (concentration 6.8×10^{-4} M) and CCl_4 (concentration 2.04×10^{-4} M) obtained after photoexcitation by UV light (285 nm for CHCl_3 and 320 nm for CCl_4 , excitation time 5 min for CHCl_3 and 10 min for CCl_4); 1, for CHCl_3 and 2, for CCl_4 .

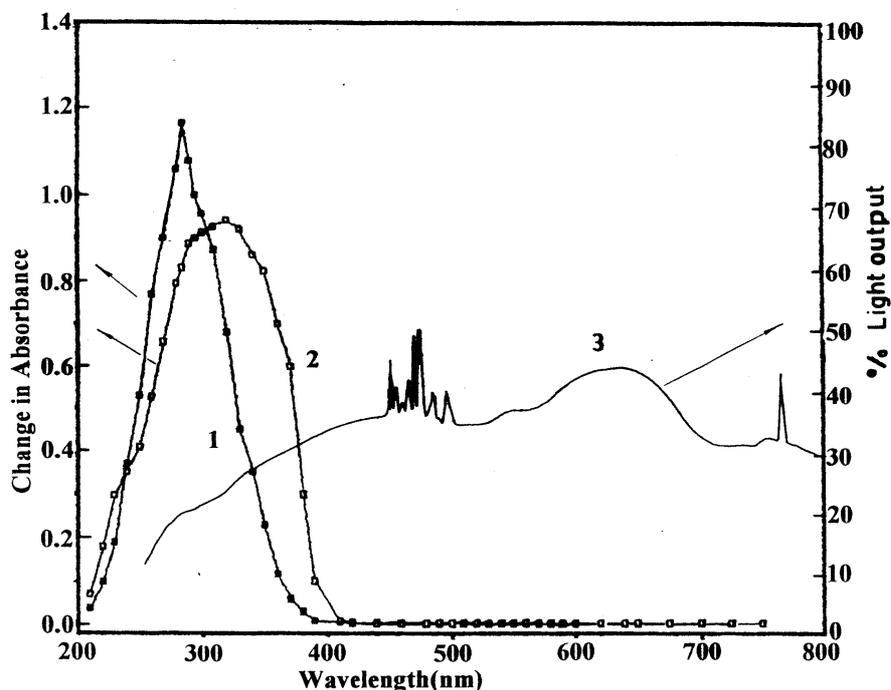


Figure 3. Plot of the change in absorbance of the band at (362 nm for CHCl_3 and 367 nm for CCl_4) versus exciting wavelength studied for solutions of ferrocene in CHCl_3 (1) and CCl_4 (2). Concentration of ferrocene 6.8×10^{-4} M. Photoexcitation for 10 min. Curve 3, intensity profile of xenon lamp

and two shoulder bands around 315 and 530 nm. The curve 2 in figure 2 for CCl_4 shows four clear absorption peaks at ≈ 290 , 320, 367 and 627 nm and one weak shoulder at ≈ 530 nm in the spectrum.

The effects of photoexcitation on the electronic absorption spectra of ferrocene in CHCl_3 and CCl_4 were studied as a function of exciting wavelength at intervals of a few nanometres in the spectral range 220–750 nm. The change in absorbance of the new bands as a function of exciting wavelength was checked and such a plot for the change in absorbance of the band at 362 nm as a function of exciting wavelength is shown in figure 3. This figure shows that the photoeffects depend drastically on the exciting wavelength. The curves 1 and 2 in figure 3 demonstrate clearly the profile of the action spectrum for the photoinduced changes in the case of ferrocene– CHCl_3 and ferrocene– CCl_4 systems, respectively. In the UV region the maximum intensity of the photoinduced changes was observed for irradiation at about 285 and 320 nm for CHCl_3 and CCl_4 , respectively. The action spectrum is broad for the ferrocene– CCl_4 interaction and the action spectrum for the ferrocene– CHCl_3 interaction is comparatively sharp. The profile of the action spectrum is possibly related to the solvent effects. Similar results were obtained for other absorption bands also. However, for both the solvents no changes in the spectra were observed for photoexcitation with visible light. The Xenon lamp intensity profile, shown in figure 3 (curve 3), helps to understand the action spectra shown in figure 3 (curves 1 and 2). In the visible region, although the light output (curve 3 in figure 3) from the

xenon lamp used for photoexcitation is significantly high, the photoinduced changes for excitation in the visible region were completely inefficient.

The changes in the spectra of ferrocene (upon photoexcitation) as a function of concentration of ferrocene in the solution as well as the time (duration) of photoexcitation were studied and observed to be systematic. Some representative results related to the changes in the spectra of ferrocene (upon photoexcitation) as a function of concentration of ferrocene in the CHCl_3 and CCl_4 solvents are discussed in the following paragraph.

The dependence of the intensity of the new bands that appeared on photoexcitation, upon the concentration of ferrocene in the halocarbon solvents was studied keeping the photo-exciting wavelength as well as duration of excitation fixed. With increasing concentration of ferrocene from lower values, a gradual increase in the intensity of all the new bands was observed as shown in figure 4a for CCl_4 (photoexcited at 320 nm for 10 min). The changes in absorbance of the band at 362 nm with the concentration of ferrocene in CHCl_3 are shown in the curve 1 of figure 4b and the changes in absorbance of the band at 367 nm with concentration of ferrocene in CCl_4 are shown in the curve 2, figure 4b. In general, figure 4b shows that for the higher values of concentrations of ferrocene the intensity of the absorption bands attained almost a saturation value for both CHCl_3 (saturation value: 6.8×10^{-4} M) and CCl_4 (saturation value: 8.8×10^{-4} M). Similar plots as in figure 4b were obtained for other new bands also. The observed change in the absorbance as a function of concentration of ferrocene is similar in nature as observed in standard plots of change in absorbance of CT bands with concentration of donor or acceptor in a standard CT complex.

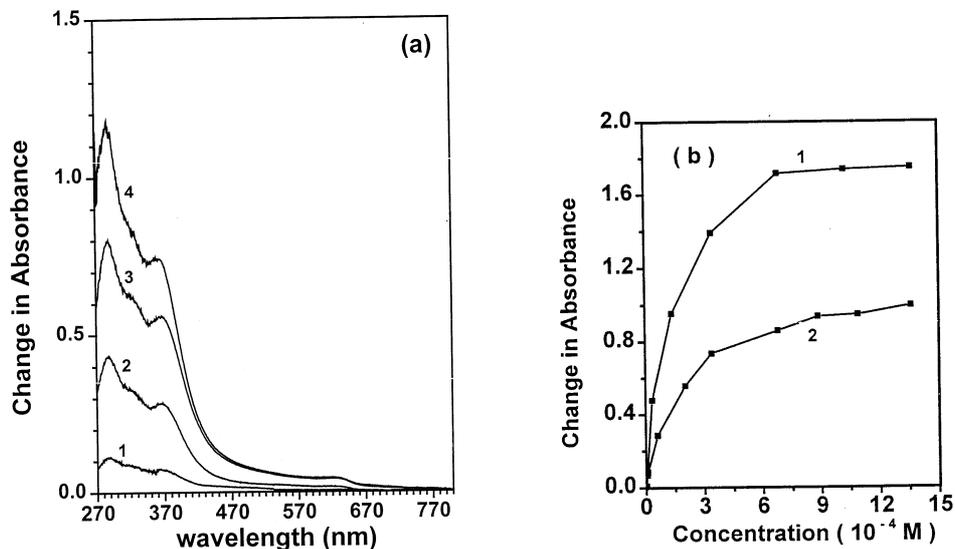


Figure 4(a). Plot of the change in absorbance versus wavelength studied for ferrocene solution in CCl_4 as a function of concentration of ferrocene using photoexcitation by 320 nm for 10 min. 1–4, refer to concentrations of 6.1×10^{-6} , 6.1×10^{-5} , 2.04×10^{-4} and 3.4×10^{-4} M, respectively. **(b)** Plot of the change in absorbance of the band (at 362 nm for CHCl_3 and 367 nm for CCl_4) versus concentration of ferrocene solutions in CHCl_3 (curve 1) and CCl_4 (curve 2) for photoexcitation (10 min) at 285 nm for CHCl_3 and at 320 nm for CCl_4 .

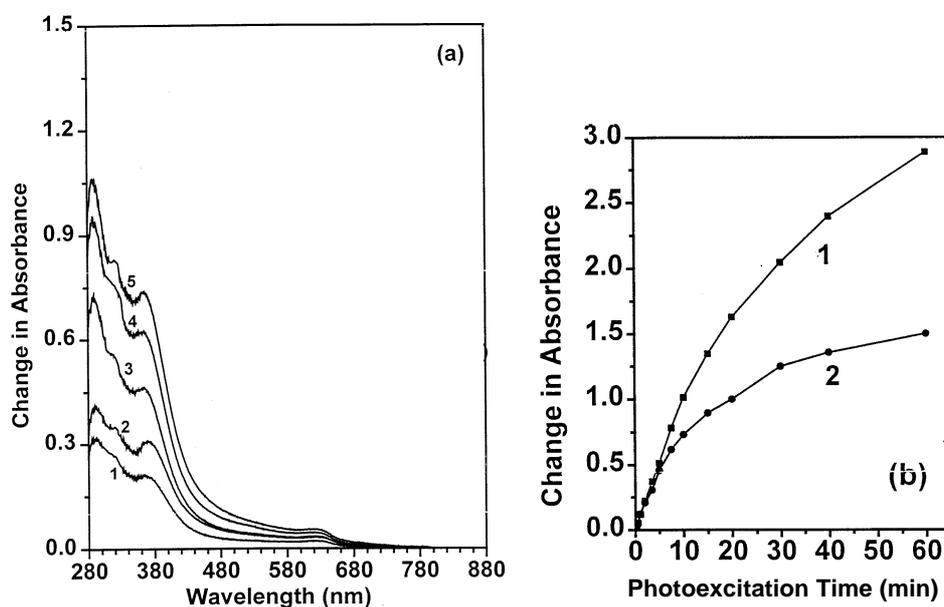


Figure 5(a). Change in the electronic absorption spectra of ferrocene in CCl₄ solution (deoxygenated; concentration: 6.8×10^{-4} M) for different time durations of photoexcitation by 320 nm. Curves 1–5 refer to photoexcitation time 2, 3.5, 5, 7.5 and 10 min respectively. **(b)** Plots of the change in absorbance of the band (at 362 nm for CHCl₃ and 367 nm for CCl₄) versus photoexcitation time. Curve 1, for ferrocene in CHCl₃ solution (deoxygenated, photoexcited by 285 nm for 10 min). Curve 2, for ferrocene in CCl₄ solution (deoxygenated, photoexcited by 320 nm for 10 min). Concentration: 6.8×10^{-4} M.

The variation in the electronic absorption spectra of ferrocene in halocarbon solvents with time (duration) of photoexcitation by UV light was studied. Enhancement in the intensity of the absorption bands with time of photoexcitation was noticed as shown in figure 5a for CCl₄ for excitation by 320 nm. Similar results related to the time of photoexcitation were obtained for CHCl₃ also. In figure 5b, plots of change in absorbance of the band at 362 nm vs. time of photoexcitation are represented for CHCl₃ by curve 1. The curve 2 of figure 5(b) represents the plot of the change in the absorbance at 367 nm vs. time of photoexcitation for CCl₄. From the curves 1 and 2 in figure 5b it is seen that the change in absorbance increases linearly with increasing time of photoexcitation up to about 10 min and for higher values of excitation time such linear relationship does not hold good. Similar plots as in figure 5b were obtained for other new bands also. The results show that the kinetics of photoinduced changes is of first order for the photoexcitation of smaller duration (up to ≈ 10 min). From the nature of the variation (as a function of time of photoexcitation) of the absorbance of the new bands in the electronic absorption spectra of ferrocene in halocarbon solvents obtained after photoexcitation, it appears that the photoinduced changes are systematic and must have some physical basis.

It has already been mentioned earlier that the new bands in the electronic absorption spectra of ferrocene (at room temperature) in the solutions of CHCl₃ after photoexcitation

appear at $\approx 290, 315, 362, 530$ and 620 nm. Similarly, in the case of CCl_4 , after photoexcitation the new bands in the spectrum appear at $\approx 290, 320, 367, 530$ and 627 nm. Scott and Becker originally reported¹⁵ a weak absorption band (shoulder) of ferrocene at about 528 nm in few solvents and they assigned this band as a spin forbidden singlet-triplet absorption corresponding to the forbidden singlet-singlet absorption at 324 nm. It should be mentioned here that other investigators have studied the absorption spectra of ferrocene at room temperature. Other investigators have not reported the shoulder at 528 nm. In the present study also the shoulder at about 528 nm has not been noticed in the spectra of ferrocene in the halocarbon solvents before photoexcitations. However, in the spectra after photoexcitations a shoulder has been observed to appear at ≈ 530 nm. The position of this band is close to the band peaked at 528 nm, assigned by Scott and Becker¹⁵ as a spin forbidden singlet-triplet absorption corresponding to the forbidden singlet-singlet absorption at 324 nm. Possibly the band peaked at 530 nm is the spin forbidden singlet-triplet absorption band of ferrocene, as assigned by Scott and Becker mentioned above. The appearance of the forbidden singlet-triplet absorption band of ferrocene peaked at 530 nm has become possible by photoexcitation. The process, by which the band appears through photoexcitation, may be termed perturbation. We do not mean the perturbation of energy levels of FcH^+ or generation of an entirely new species. But we believe that as a result of photoexcitation the forbidden singlet-triplet absorption band of ferrocene has become 'allowed'. Solutions of ferrocene in CHCl_3 and CCl_4 have shown other new absorption maxima at $\approx 290, 362$ and 620 – 627 nm. The band at 620 – 627 nm is due to ferricenium ion¹² and the bands at 290 and 362 nm appear due to organic radical (R) cations formed¹⁰⁻¹² as a result of photoreaction. The formation of ferricenium ion [FcH^+] and radical cation indicate that the new band appeared at $315/320$ nm in halocarbon solvents $\text{CHCl}_3/\text{CCl}_4$ after photoexcitations originates due to ferrocene-to-solvent-charge-transfer (CTTS) transitions. The CT band overlaps with the bands of ions as well as ferrocene, which make this band broader (shoulder type).

It has been discussed earlier that in the UV region, the maximum intensity of the photoinduced changes was observed for irradiation at about 285 and 320 nm for ferrocene- CHCl_3 and ferrocene- CCl_4 systems, respectively. It is interesting to note that in the case of ferrocene- CCl_4 system the maximum intensity of the photoinduced changes was noticed for irradiation at the wavelength corresponding to the position of the CTTS transition. But for the ferrocene- CHCl_3 system the maximum intensity of the photoinduced changes was observed for irradiation at 285 nm while the position of the CTTS transition was at 315 nm. The overall profile of the action spectrum for the photoinduced changes is possibly related to the solvent effect.

The primary step in the photoprocess is the dissociation of the CTTS state to give ferricenium cation [FcH^+], Cl^- and organic radical ion (CHCl_2 , in the case of CHCl_3 and CCl_3 , in the case of CCl_4)³. On the basis of the results one could consider the following primary photochemical process:



for the formation of ferricenium ion and this process occurs through CTTS excitation. In fact, the real mechanism of the photoinduced processes involved is much complicated. It is not possible to monitor this from the present spectral studies.

Considering Mulliken's theory for formation of a charge-transfer complex, to a first approximation one can relate the energy of the charge-transfer band with the ionization

potential of the donor and the electron affinity of the acceptor molecule by the linear relation¹⁶:

$$h\nu_{\text{CT}} = I_{\text{D}}^{\text{v}} - E_{\text{A}}^{\text{v}} + C_1 \quad (2)$$

where, $h\nu_{\text{CT}}$ is the energy of the lowest-energy intermolecular charge-transfer band, I_{D}^{v} is the vertical ionization potential of the donor, E_{A}^{v} is the vertical electron affinity of the acceptor and C_1 is a constant. If charge-transfer complexes of different donor molecules are formed with a common acceptor molecule, according to (2) a plot of $h\nu_{\text{CT}}$ against I_{D}^{v} is expected to be linear with a slope of unity. Unfortunately, reliable values of vertical ionization potential and electron affinities are very scarce. Apart from the necessary distinction between the vertical and adiabatic or absolute values, different experimental methods yield different results, and also sometimes the substance of interest has not been studied at all. Usually for the practical purposes, the ionization potentials are considered as I_{D} .

The formation of charge-transfer complexes of different electron donors with CCl_4 has been reported in the literature^{17,18}. For a ready reference the positions of the charge-transfer band observed for the CT complexes of CCl_4 with different donor molecules are shown in table 1. The values of I_{D} of the donors including the references are also presented in this table 1. A plot of $h\nu_{\text{CT}}$ against I_{D} for the CT complexes of CCl_4 with different donor molecules is shown in figure 6. A satisfactory straight line as predicted by (2) is obtained. The use of charge-transfer spectra for the evaluation of ionization potentials is well known. Now, if the new band at 320 nm observed for the solution of ferrocene in CCl_4 after photoexcitation arises due to the formation of CT complex of ferrocene with the solvent CCl_4 , one can estimate the value of I_{D} of ferrocene with the help of the linear plot in figure 6 considering the value (3.87 eV) of $h\nu_{\text{CT}}$ corresponding to the band at 320 nm. The value of I_{D} estimated for ferrocene from figure 6 is 6.05 eV. The experimentally measured (photoionization threshold) value of I_{D} of ferrocene is reported²³ to be 6.1 eV. The excellent agreement between the values of estimated and experimental I_{D} of ferrocene supports the charge-transfer concept for the complex between ferrocene and CCl_4 (after photoexcitation) under study and also supports the new band at 320 nm observed for the solution of ferrocene in CCl_4 after photoexcitation to be the CTTS band. From (2) one expects a slope of unity for the $h\nu_{\text{CT}}$ – versus I_{D} plot. The linear plot in figure 6 gives a slope of 0.60. This low value for the slope could be due

Table 1. Position of the charge-transfer (CT) band observed for the CT complexes of CCl_4 with different donor molecules.

Donor	Position of CT band (nm)	Value of $h\nu_{\text{CT}}$ (eV)	Value of I_{D} (eV)
Benzene	216 ¹⁷	5.74	9.24 ¹⁹
<i>m</i> -Xylene	230 ¹⁷	5.39	8.59 ¹⁹
<i>p</i> -Xylene	230 ¹⁷	5.39	8.44 ¹⁹
Mesitylene	235 ¹⁷	5.28	8.39 ¹⁹
Hexamethylbenzene	242 ¹⁷	5.15	8.00 ^{20,21}
Tetracene	285 ^{#18}	4.35	6.94 ²²

[#]Value estimated from ref. 18

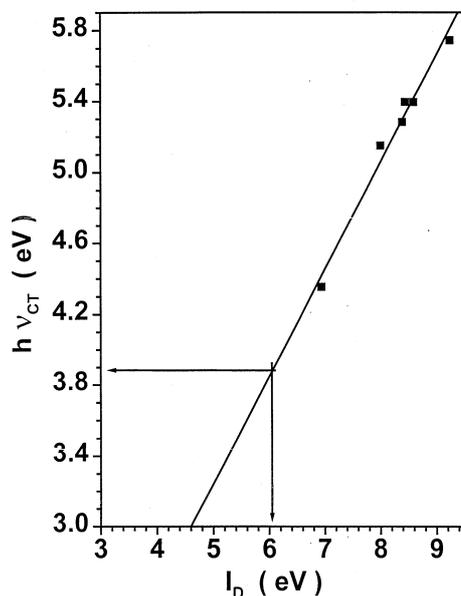


Figure 6. Plot of $h\nu_{CT}$ versus I_D for the charge-transfer (CT) complexes of different electron donors with CCl_4 as an electron acceptor. The horizontal and vertical arrows indicate the position of the CT band in the case of the CT complex of ferrocene with CCl_4 and the value of estimated I_D for ferrocene, respectively.

to the fact that ionization potential values used for the plot were not vertical values. Further, (2) is only an approximate one. Indeed, such a deviation of slope from unity is a general observation in these types of experiments^{24–26}. McConnell *et al*²⁵ have mentioned that the slope may have any value between 0.4–1.0. It has been indicated by Foster²⁶ that a large deviation from straight line of unit gradient in a plot like (2) can appear if the electron donor is only weakly effective. In the case of all-trans-*b*-carotene, the plot of CT band energy against electron affinity of some acceptor molecules was reported²⁴ to be linear with a slope of 0.44. But the ionization potential evaluated (5.44 eV) for all-trans-*b*-carotene agreed²⁴ well with the experimental value estimated (5.5 eV) from photoemission threshold. Thus, the matching of the calculated ionization potential of ferrocene with experimental value is not accidental. If the band at 307 nm for the solution of ferrocene in CCl_4 is considered as CTTS band as suggested by the previous research workers, the estimated value of I_D for ferrocene from figure 6 appears to be 6.32 eV, which cannot give better agreement with the experimental value (6.1 eV)²³. Thus, the excellent agreement between the estimated value and the experimentally measured value of ionization potential indicates the exactness of the located band position for the CTTS transition in the case of the solution of ferrocene in CCl_4 . It has been mentioned earlier that the new band appeared at 315 nm in the case of CHCl_3 after photoexcitation originates due to ferrocene-to-solvent charge-transfer (CTTS) transition. This data could not be utilized for identifying the CTTS position for the ferrocene- CHCl_3 interaction (as in figure 6 in the case of ferrocene- CCl_4 interaction) because the data available for the reported well resolved charge-transfer band for the CT complexes of CHCl_3 with different donor molecules are inadequate.

4. Conclusion

Changes in the electronic absorption spectra of ferrocene in the solutions of halocarbon solvents (CHCl_3 and CCl_4) have been observed under photoexcitations by UV light in the nitrogen atmosphere. Using Mulliken's theory for the formation of a charge-transfer complex and considering the new band position (320 nm) for CTTS transition in the case of carbontetrachloride, ionization potential of ferrocene has been estimated. The estimated value has shown excellent agreement with the experimental value. Experimental results show that the new band appeared at ≈ 315 and 320 nm in the electronic absorption spectrum of ferrocene in the solutions of CHCl_3 and CCl_4 , respectively, after photoexcitation originates due to the CTTS transition under photoexcitation.

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References

1. Carraher C E, Sheets J E and Pittman C U 1982 *Advances in organometallic and inorganic polymer science* (New York: Marcel Dekker)
2. Turner A P F, Karube I and Wilson G S (eds) 1987 *Biosensors* (London: Oxford Univ. Press)
3. Geoffroy G L and Wrighton M S 1979 *Organometallic photochemistry* (New York: Academic Press) p. 242
4. Meier H 1974 *Organic semiconductors* (Weinheim: Verlag Chemie) p. 165
5. Gutmann F, Keyzer H and Lyons L E 1983 *Organic semiconductors* (Malabar, FL: Krieger) part B, pp. 360, 463
6. Ozaki J, Watanabe T and Nishiyoma Y 1993 *J. Phys. Chem.* **97** 1400
7. Mallik B and Chakraborty A K 1996 *J. Phys. Chem.* **100** 2145
8. Mallik B and Chakraborty A K 1997 *J. Chem. Soc., Faraday Trans.* **93** 3677
9. Chakraborty A K and Mallik B 1995 *Synth. Met.* **73** 239
10. Akiyama T, Sugimori A and Hermann H 1973 *Bull. Chem. Soc. Jpn.* **46** 1855
11. Traverso O and Scandola F 1970 *Inorg. Chim. Acta* **4** 493
12. Brand J C D and Snedden W 1957 *Trans. Faraday Soc.* **53** 894
13. Augustyniak A J and Wojtezak J 1984 *Trans. Metal Chem.* **9** 303
14. Tatistcheff H B, Hancock L F and Wrighton M S 1995 *J. Phys. Chem.* **99** 7689
15. Scott D R and Becker R S 1961 *J. Chem. Phys.* **35** 516
16. Chen E C M and Wentworth W E 1975 *J. Chem. Phys.* **63** 3183
17. Kellawi H and Rosseinsky D R 1969 *J. Chem. Soc. (A)* 1207
18. Davis K M C and Farmer M F 1968 *J. Chem. Soc. (B)* 859.
19. Gutmann F and Lyons L F 1967 *Organic semiconductors* (New York: John Wiley) p. 680
20. Meyer F and Harrison A G 1964 *Can. J. Chem.* **42** 2256
21. Foster R 1969 *Organic charge-transfer complexes* (London and New York: Academic Press) p. 384
22. Gutmann F, Keyzer H and Lyons L E 1983 *Organic semiconductors* (Malabar, FL: Krieger) part B, p. 497
23. Gutmann F, Keyzer H and Lyons L E 1983 *Organic semiconductors* (Malabar, FL: Krieger) part B, p. 501
24. Mallik B, Jain K M and Misra T N 1980 *Biochem. J.* **189** 547
25. McConnell H, Ham J S and Platt J R 1953 *J. Chem. Phys.* **21** 66
26. Foster R 1960 *Tetrahedron* **10** 96