

# Bioactive nanocrystalline wollastonite synthesized by sol–gel combustion method by using eggshell waste as calcium source

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**Abstract.** The sol–gel combustion method was employed to synthesize the nanocrystalline wollastonite by taking the raw eggshell powder as a calcium source and TEOS as a source of silicate. Glycine was used as a reductant or fuel and nitrate ions present in metal nitrate acts as an oxidizer. The phase purity of the wollastonite was analysed by powder XRD and the product is found to contain single-phasic wollastonite. FT–IR spectrum shows the characteristic peaks of the functional groups present in the wollastonite. SEM images show that particles are agglomerated and the particle size is found to be in the nanoregime. The calcination temperature was optimized based on the thermal analysis of the precursor. The bioactivity of wollastonite was investigated by immersing the scaffold in a simulated body fluid for 15 days at 37 °C and intermediate analysis of the surface by XRD shows the deposition of hydroxyapatite layer after 5 days.

**Keywords.** Bioceramic; raw eggshell powder; sol–gel combustion; bioactivity; hydroxyapatite.

## 1. Introduction

Ceramics like wollastonite, hydroxyapatite (HA),  $\beta$ -tricalcium phosphate are the various bioceramics have shown to bond with living bone directly hence, they received important attention in the field of biomedical engineering as a significant bone substitute (Sainz *et al* 2010). The characteristic of a bioactive material is its ability to form hydroxyapatite layer on its surface in presence of physiological environment, which provides the bonding interface with the tissues and bone (Wei *et al* 2009). Calcium silicate ceramics were used as an alternate for hydroxyapatite in the biomedical applications like artificial dental roots and as a bone regenerative material (Wang *et al* 2008). Calcium silicates are also used in industrial, technological applications due to its properties like high temperature strength, creep resistance, chemical inertness, thermal stability, low thermal expansion and low thermal conductivity (Sreekanth Chakradhar *et al* 2006).

Wollastonite belongs to the group of bioactive and biocompatible calcium silicates. In recent years wollastonite has received much attention due to its high bioactivity and biocompatibility which influenced the researchers to study its application in the field of drug delivery system. The bioactivity of wollastonite is higher than any of the other biocompatible glass and glass ceramics. The bioactivity of calcium silicate was analysed by the rate of hydroxyapatite formation on its surface in presence of simulated body fluid (SBF) (Huang and Chang 2009). The bioactivity of wollastonite is

due to silicon ion present in it, as it plays an essential role in the formation of new bone by metabolic process. Wollastonite shows an excellent bone-bonding ability than the hydroxyapatite and tricalcium phosphate (Kobayashi *et al* 1998).

The preparation of wollastonite was reported for hydrothermal method (Lin *et al* 2006), sol–gel method (Meiszterics and Sinko 2008), solution combustion method (Sreekanth Chakradhar *et al* 2006), co-precipitation (Binnaz Yoruc and Hazar 2007), citrate–nitrate gel combustion method (Huang and Chang 2009) and microwave synthesis (Vichaphund *et al* 2011). The preparation procedure plays a vital role in the bioactivity of the calcium silicates as the sol–gel synthesized wollastonite is found to have higher bioactivity than the wollastonite prepared by solid-state method. However, the agglomerates produced were homogeneous in sol–gel synthesis than the products synthesized by any of the above mentioned methods (Ortega–Lara *et al* 2010). The phase transformation temperature of the solution combustion synthesized wollastonite powder was found to be lower than the powders obtained by solid-state method (Sreekanth Chakradhar *et al* 2006).

The sol–gel combustion process is a combination of sol–gel and combustion syntheses. In sol–gel combustion process the vigorous redox reaction between the fuel and oxidant is highly exothermic and releases large amount of heat energy at which the phase formation is completed for most of the compounds hence the calcination temperature is found to be lower for this method when compared to the sol–gel method, combustion method and solid-state method. Wollastonite synthesized by microwave method by using biowaste and eggshell as a source of calcium is found to possess a good

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bioactivity (Vichaphund *et al* 2011). Wollastonite synthesized by the pyrolysis of chicken eggshells followed by sol-gel process and the product was reported to possess dielectric property and it is used in the field of biomedical engineering (Tangboriboon *et al* 2011).

In the present work, we are reporting the synthesis of wollastonite by sol-gel combustion method, by using glycine as a fuel and raw eggshell powder as a calcium source. The product formed was analysed for its *in vitro* bioactivity in the simulated body fluid medium.

## 2. Experimental

### 2.1 Materials

Chicken eggshell, glycine (Qualigens fine 98%), tetraethyl orthosilicate (Acros organics 98%), concentrated nitric acid, hydrochloric acid, sodium chloride (s.d. fine 99.9%), sodium bicarbonate (nice chemical 99%), potassium chloride (s.d. fine 99.5%), dipotassium hydrogen phosphate (s.d. fine 99%), calcium chloride (Qualigen fine 98%), magnesium chloride (Ranbaxy 97%), sodium sulphate (s.d. fine 99.5%) and tris hydroxymethylate aminomethane (s.d. fine 99%) were used.

### 2.2 Experimental procedures

The chicken eggs were collected from Vellore market and boiled in beaker containing water for 20 min on Bunsen burner. The eggshells were removed and dried in hot air oven at 120 °C for 2 h. The raw eggshell is grinded in mortar and pestle for 30 min to get fine powders.

The fine powder of eggshell was reported to contain 94.37% CaCO<sub>3</sub>, hence in order to prepare 1 M Ca<sup>2+</sup> ion solution 10 g of raw eggshell was taken and dissolved in 5 ml con. HNO<sub>3</sub>. The unreacted constituents of eggshell were filtered by Watmann 41 filter paper and the resultant solution was made up to 100 ml. This solution with approximately 1 M concentration of Ca<sup>2+</sup> ions is used as a stock solution for further experiments.

### 2.3 Estimation of Ca<sup>2+</sup> ions in eggshell solution

The concentration of calcium ions present in the raw eggshell solution was estimated by complexometric titration and flame photometry. In flame photometry CaCl<sub>2</sub> (0.277 g/100 ml) was taken as a standard solution to plot the standard graph and the stock solution prepared from eggshell was sufficiently diluted and analysed for the concentration of Ca<sup>2+</sup> ions.

In the complexometric titration, CaCO<sub>3</sub> (0.1 g/100 ml) dissolved in HCl and made up to 100 ml was taken as a standard solution and EDTA was taken as a complexing agent. From the standard solution the concentration of calcium present in eggshell solution was estimated after sufficient dilution and it was found to be 0.01 M.

### 2.4 Synthesis of wollastonite

The stoichiometric ratios of eggshell solution and glycine solution were taken and mixed thoroughly. Tetraethyl orthosilicate (TEOS 1 M) was added to the above mixture with constant stirring for 10 min at room temperature. The rate of hydrolysis of TEOS is slow in neutral solutions and requires an acid or base catalyst hence the pH of the solution was adjusted to 1 by adding concentrated nitric acid. The resultant solution was stirred constantly and the temperature was raised to 40 °C and maintained for 90 min in a covered glass beaker. The thick gel formed was dried at 100 °C for 7 h and the decomposition of the gel was carried out at 400 °C for 2 h in muffle furnace. The precursor formed was calcined at 900 °C for 12 h results in a white product which was analysed by powder XRD in order to check the phase purity.

### 2.5 Preparation of scaffold

200 mg of the calcined wollastonite powder was made into scaffold by applying 7 tons of pressure in hydraulic pellet press with the scaffold diameter of 12 and 3 mm thickness for the bioactivity studies.

### 2.6 *In vitro* bioactivity assessment of scaffold in simulated body fluid

Simulated body fluid (SBF) was prepared by other procedures by dissolving NaCl, NaHCO<sub>3</sub>, KCl, K<sub>2</sub>HPO<sub>4</sub>·3H<sub>2</sub>O, MgCl<sub>2</sub>·6H<sub>2</sub>O, CaCl<sub>2</sub>, Na<sub>2</sub>SO<sub>4</sub>, (CH<sub>2</sub>OH)(NH<sub>2</sub>) in deionized water and HCl solution was added to adjust the pH 7.25. The prepared simulated body fluid has the ionic concentration similar to human blood plasma (Siriphannona *et al* 2002; Wei *et al* 2009). The *in vitro* bioactivity of wollastonite scaffold was investigated by placing the scaffold in 15 ml simulated body fluid (SBF) kept at 37 °C in an incubator with the replacement of the simulated body fluid at the end of every 24 h. The scaffold was removed at various time periods and analysed for the growth of hydroxyapatite (HAp) layer on the surface of the scaffold (Raquez *et al* 2011). The scaffold was dried at 80 °C for 15 min in a hot air oven, the formation of the hydroxyapatite layer on the surface of wollastonite was analysed by powder XRD on 5<sup>th</sup>, 10<sup>th</sup> and 15<sup>th</sup> day of incubation.

### 2.7 Characterization

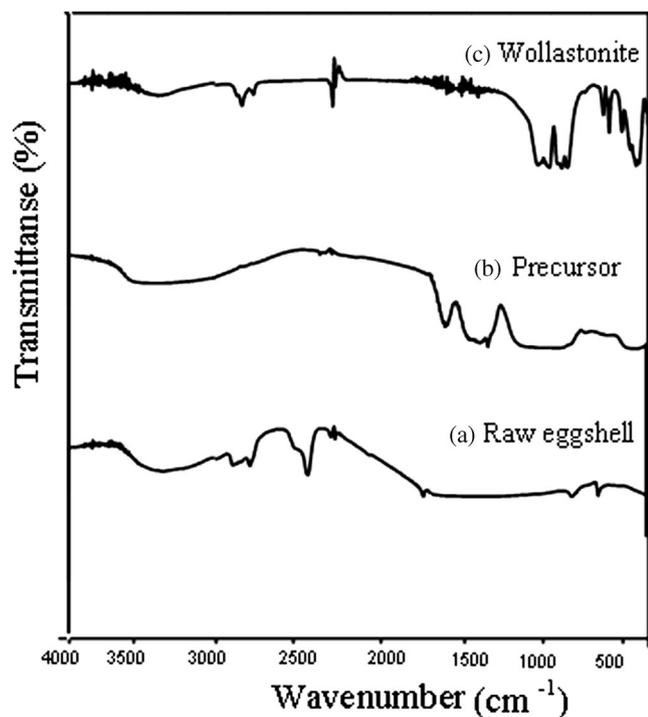
Phase identification was done by using Bruker D8 Advance X-ray diffractometer, using CuKα (1.5406 Å) and Ni filtered radiation. The functional groups were identified using FT-IR spectrometer (Shimadzu). SEM images were recorded by using Hitachi S-3400N Model, at an accelerating voltage of 15 kV. The thermodynamic property of precursor was studied by the thermo-gravimetric analyser (TGA) and differential thermal analyser (DTA). *In vitro* bioactivity studies were carried out in orbital shaking incubator RIS-24B.

### 3. Results and discussion

When eggshell is dissolved in concentrated nitric acid,  $\text{CaCO}_3$  ions present in the eggshell will get converted to  $\text{Ca}(\text{NO}_3)_2$ . pH is adjusted to 1 by adding concentrated nitric acid, the tetraethyl orthosilicate in acidic medium will undergo hydrolysis and results in the formation of alcohol and  $\text{Si}(\text{OH})_4$ . The resultant  $\text{Si}(\text{OH})_4$  reacts with  $\text{Ca}^{2+}$  ions present in the eggshell solution to form amorphous  $\text{CaSiO}_3$ . Glycine plays a major role as a fuel for the combustion. When the mixture was stirred, hydrolysis of TEOS results in the formation of ethyl alcohol which undergoes condensation reaction with the glycine and forms a polymeric network. The nitrate ions from eggshell solution and nitric acid added for hydrolysis will act as an oxidizer during the decomposition of gel and results in an exothermic reaction. The high temperature evolved during combustion aids in the phase formation of wollastonite.

#### 3.1 Fourier transform infrared spectroscopy

FTIR spectrum was recorded in the range of  $400\text{--}4000\text{ cm}^{-1}$ , with the resolution of  $4\text{ cm}^{-1}$ . FT-IR spectrum (figure 1a) of raw eggshell show peaks at  $712, 874, 1459$  and  $2519\text{ cm}^{-1}$  which are the characteristic peaks of  $\text{CaCO}_3$ , the absorption that occurred around  $1798, 2880$  and  $2982\text{ cm}^{-1}$  correspond to  $\text{C}=\text{O}$  from carbonate ion ( $\text{CO}_3^{2-}$ ). The very strong band at  $1459\text{ cm}^{-1}$  is related to the carbonate  $\text{C}-\text{O}$  bonds, band at around  $588\text{ cm}^{-1}$  corresponds to  $\text{Ca}-\text{O}$  bonds. And the broad band at  $3413\text{ cm}^{-1}$  corresponds to the moisture



**Figure 1.** FTIR spectra of the (a) raw eggshell powder, (b) precursor and (c) wollastonite.

absorption on the surface of the pellet (Wang *et al* 2006; Miguel Galván-Ruiz *et al* 2009). Figure 1(b) shows peaks at  $470\text{--}1023\text{ cm}^{-1}$  due to  $\text{Si}-\text{O}$ ,  $\text{O}-\text{Si}-\text{O}$ ,  $\text{Si}-\text{O}-\text{Ca}$ ,  $\text{Si}-\text{O}-\text{Si}$  groups and  $1471\text{--}1641\text{ cm}^{-1}$  due to stretching vibration of carbonate ions.

Figure 1(c) shows bending modes of  $\text{Si}-\text{O}-\text{Si}$  and  $\text{O}-\text{Si}-\text{O}$  bonds at  $476\text{ cm}^{-1}$ , stretching modes of  $\text{O}-\text{Si}-\text{O}$  bonds at  $798\text{ cm}^{-1}$ ,  $\text{Si}-\text{O}-\text{Ca}$  bonds containing non-bridging oxygen at  $910\text{--}944\text{ cm}^{-1}$ . The symmetric stretching vibrations of  $\text{Si}-\text{O}-\text{Si}$  bonds were observed at  $1022\text{--}1083\text{ cm}^{-1}$  and bending vibration of  $\text{H}_2\text{O}$  at  $1630\text{ cm}^{-1}$ . The band at  $3437\text{ cm}^{-1}$  is due to the absorption of moisture on the surface of the wollastonite. The carbonate ions present in traces in the precursor is completely removed from the product by calcining at  $900\text{ }^\circ\text{C}$  for 12 h.

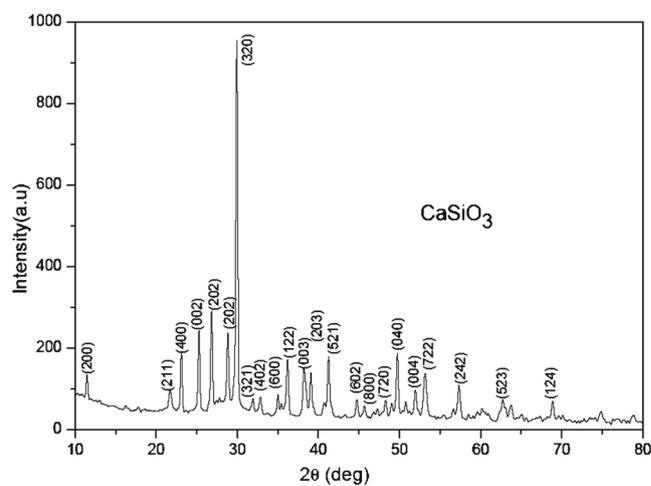
#### 3.2 X-ray diffraction

Figure 2 shows XRD pattern of the precursor calcined at  $900\text{ }^\circ\text{C}$  for 12 h and the peaks in the pattern is indexed to wollastonite (JCPDS file no: 043-1460). XRD pattern shows a single phasic wollastonite and no other secondary phases or unreacted starting materials were observed. The space group of wollastonite is  $P21/a$  (14) and crystal symmetry is monoclinic respectively. The lattice parameter calculated from the powder XRD were found to be  $a = 15.463$ ,  $b = 7.342$  and  $c = 7.193\text{ \AA}$ .

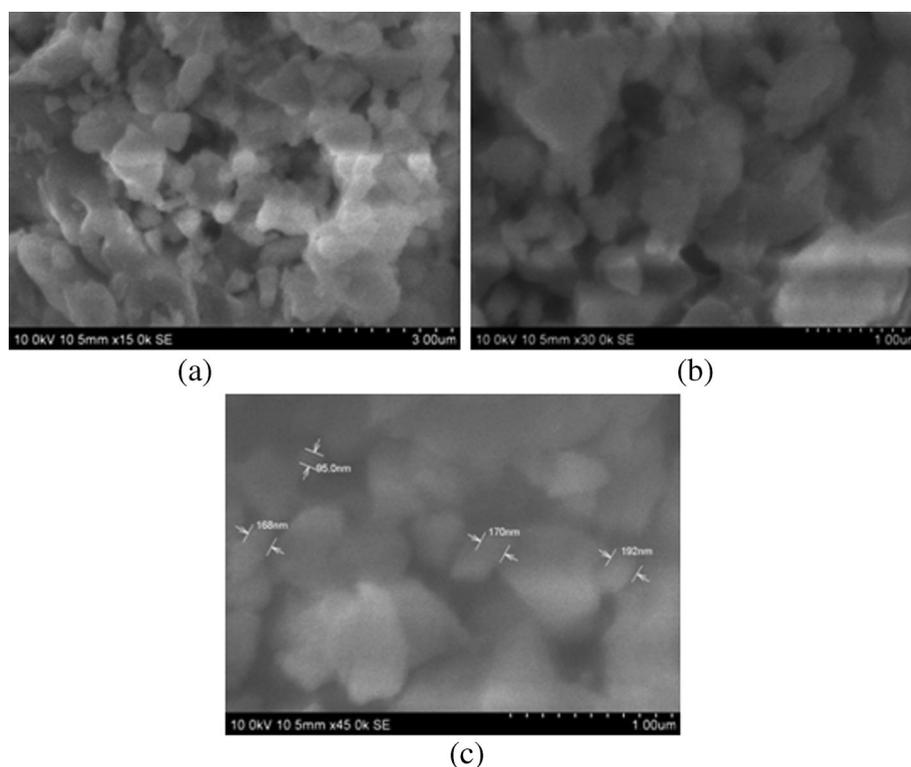
The average crystallite size ( $D$ ) of wollastonite was estimated from the full width at half-maximum (FWHM) of the powder XRD pattern using Scherrer's formula (Klug and Alexander 1962).

$$D = k\lambda / B \cos \theta,$$

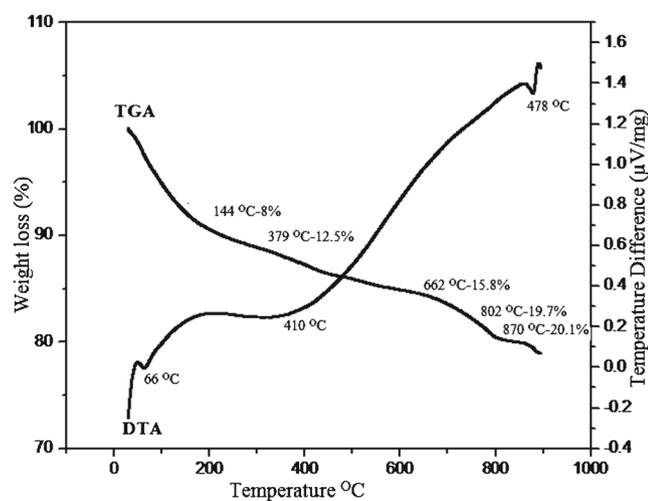
where  $\lambda$  is the wavelength of X-ray,  $\theta$  the Bragg angle and  $k$  is the constant depends on the grain shape ( $0.89\text{ \AA}$  for circular grains). The calculated result shows the average crystallite size of the calcined  $\text{CaSiO}_3$  samples in the range of  $42.0\text{ nm}$ .



**Figure 2.** XRD pattern of the wollastonite synthesized by using eggshell as a calcium source.



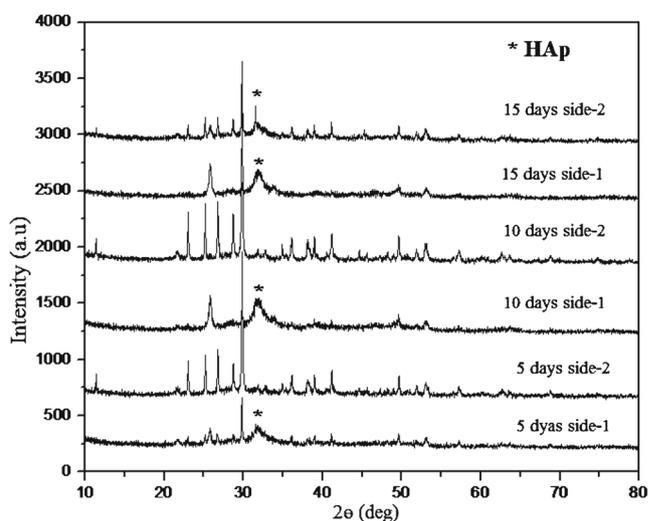
**Figure 3.** SEM images of the calcined wollastonite powder at 900 °C for 12 h.



**Figure 4.** TGA and DTA curves for the precursor.

### 3.3 Scanning electron microscopy

SEM micrographs (figure 3) of the calcined wollastonite powders show that the synthesized wollastonite powders possess irregular morphology. Figure 3(c) shows that particle size of wollastonite powder is in the nano-regime and the particles are highly heterogenous. The particle size is in the range of 95–192 nm, which is in agreement with the particle size values calculated from XRD pattern by using Scherrer's equation. Figure 3(a and b) shows that powder is highly

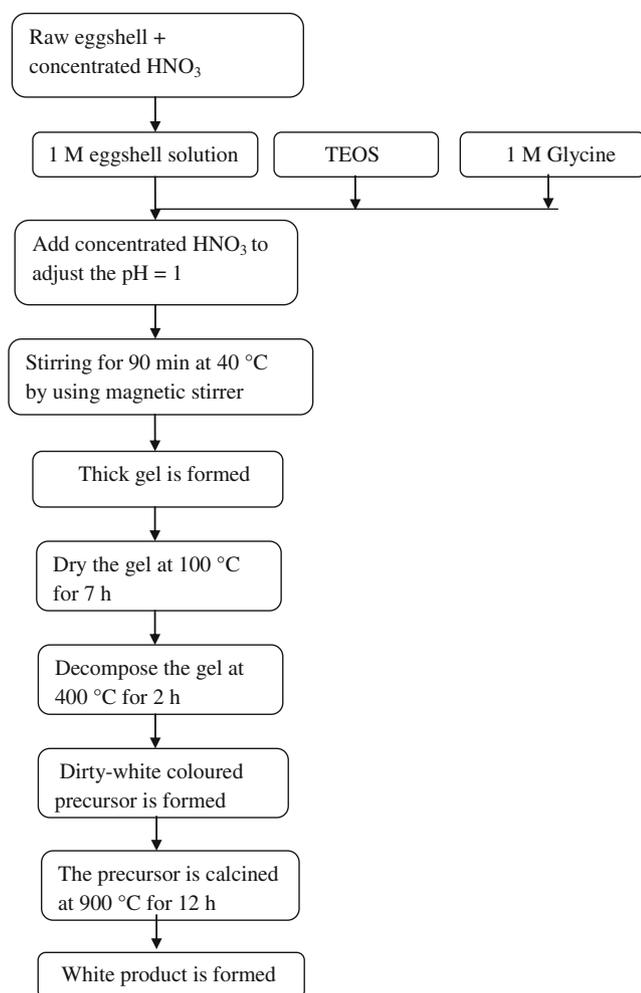


**Figure 5.** XRD pattern of the HAp deposit on the wollastonite for 5, 10 and 15 days.

agglomerated with high porosity which will result in high surface area.

### 3.4 Thermal analysis

Thermal decomposition (figure 4) of the precursor was examined by the thermo gravimetric analyser (TGA) and



Flow chart for the synthesis of nanocrystalline wollastonite by sol-gel combustion method.

differential thermal analyser (DTA) with the heating rate of  $10\text{ }^{\circ}\text{C min}^{-1}$  in the nitrogen atmosphere. TGA curve shows 8% weight loss at  $144\text{ }^{\circ}\text{C}$  is due to the evaporation of moisture from the precursor, 12.5% weight loss at  $379\text{ }^{\circ}\text{C}$  may be due to the loss of nitrates and some glycine moieties, 17.5% weight loss at  $600\text{--}802\text{ }^{\circ}\text{C}$  may be due to the removal of carbon particles and the decomposition of  $\text{CaCO}_3$  present in the mixture. DTA curve shows sharp endothermic peak at  $66\text{ }^{\circ}\text{C}$  due to the loss of water molecules, the broad endothermic peak observed at  $410\text{ }^{\circ}\text{C}$  is due to the loss of nitrates and the sharp endothermic peak at  $878\text{ }^{\circ}\text{C}$  is due to the phase transformation (Wang *et al* 2008). From TGA-DTA, the gradual phase formation was observed at  $870\text{ }^{\circ}\text{C}$ , hence the calcination temperature is optimized at  $900\text{ }^{\circ}\text{C}$ .

### 3.5 Analysis of HAp formation on wollastonite pellets

The *in vitro* bioactivity of the wollastonite scaffold was studied by placing the scaffold in the simulated body fluid. The scaffold was immersed in SBF solution for 15 days in an incubator at  $37\text{ }^{\circ}\text{C}$  with the change of SBF after every 24 h

and the phase composition of the scaffold surface was examined by XRD in order to check the deposition of hydroxyapatite. When  $\text{CaSiO}_3$  scaffold is immersed in SBF, calcium and silicon ions get released from its structure and reacts with the other constituents present in the simulated body fluid. Calcium ions released from the surface of wollastonite coating is found to increase the ionic activity of the apatite formation in SBF leading to the formation of hydroxyapatite on the coated surface (Xuanyong Liu *et al* 2004; Paluszkievicz *et al* 2008).

XRD pattern (figure 5) shows formation of hydroxyapatite on the surface of wollastonite at the end of 5<sup>th</sup>, 10<sup>th</sup> and 15<sup>th</sup> day of incubation. Fifth and tenth days XRD pattern of side 1 shows more deposition when compared with side 2 which may be due to the more circulation of SBF on side 1 as it is exposed to SBF. Side 2 is kept over the glass beaker hence the circulation of SBF is very poor and as a result the deposition of hydroxyapatite is found to be very less. The scaffold was kept for 5 more days by reversing the sides, the analysis on 15<sup>th</sup> day shows the deposition of apatite layer on both the sides which indicates that the circulation of SBF on the surface of the scaffold favours the deposition of hydroxyapatite layer.

## 4. Conclusions

Single phasic wollastonite ( $\text{CaSiO}_3$ ) was successfully synthesized by sol-gel combustion method by using glycine as a fuel and raw eggshell as a source of calcium. The phase purity of the product was found to be very high and the particle size was found to be in the nano regime. The scaffold prepared from the wollastonite shows a very good bioactivity on the surface where the circulation of SBF is high and it shows poor or no activity on the surface where circulation of SBF is low. From the reports, we conclude that the method is effective in converting a biowaste in to a useful bioceramics.

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## References

- Binnaz Yoruc A and Hazar 2007 *J. Ceram. Int.* **33** 687
- Chengyu Wang, Ping Xiao, Jingzhe Zhao, Xu Zhao, Yanhua Liu and Zichen Wang 2006 *Powder Technol.* **170** 31
- Huanping Wang, Qilong Zhang, Hui Yang and Huiping Sun 2008 *Ceram. Int.* **34** 1405
- Jie Wei, Fangping Chen, Jung-Woog Shin, Hua Hong, Chenglong Dai, Jiancan Su and Changsheng Liu 2009 *Biomaterials* **30** 1080
- Kaili Lin, Jiang Chang and Jianxi Lu 2006 *J. Mater. Lett.* **60** 3007

- Klug H and Alexander L 1962 *X-ray diffraction procedures* (New York: Wiley) p. 491
- Kobayashi M, Nakamura T, Okada Y, Fukumoto A, Furukawa T, Kato H, Kokubo T and Kikutani T 1998 *J. Biomed. Mater. Res.* **42** 223
- Meiszterics A and Sinko K 2008 *Colloids Surf.* **A319** 143
- Miguel Galván-Ruiz, Juan Hernández, Leticia Baños, Joaquín Noriega-Montes and Mario E Rodríguez-García 2009 *J. Mater. Civ. Eng.* **21** 694
- Nuchnapa Tangboriboon, Tunchanoke Khongnakhon, Supawinee Kittikul Ruksapong Kunanuruksapong and Anuvat Sirivat 2011 *J. Sol-Gel Sci. Technol.* **58** 33
- Ortega-Lara W, Cortes-Hernandez D A, Best S, Brooks R and Hernandez-Ramirez A 2010 *Ceram. Int.* **36** 513
- Paluszkiewicz C, Blazewicz M, Podporska J and Gumuła T 2008 *Vib. Spectrosc.* **48** 263
- Punnama Siriphannona, Yoshikazu Kameshima, Atsuo Yasumori, Kiyoshi Okada and Shigeo Hayashi 2002 *J. Eur. Ceram. Soc.* **22** 511
- Raquez J M, Barone D T J, Luklinska Z, Persenaire O, Belayew A, Eyckmans J, Schrooten J and Dubois Ph 2011 *J. Biomacro.* **12** 692
- Sainz M A, Pena P, Serena S and Caballero A 2010 *Acta Biomater.* **6** 2797
- Sreekanth Chakradhar R P, Nagabhushana B M, Chandrappa G T, Ramesh K P and Rao J L 2006 *Mater. Chem. Phys.* **95** 169
- Vichaphund S, Kitiwan M, Atong D and Thavorniti P 2011 *J. Eur. Ceram. Soc.* **31** 2435
- Xiang-Hui Huang and Jiang Chang 2009 *Mater. Chem. Phys.* **115** 1
- Xuanyong Liu, Chuanxian Ding and Chu Paul K 2004 *Biomaterials* **25** 1755